

8. Consider the following redox reaction: $2\text{MnO}_4^- + 3\text{ClO}_2 + \text{H}_2\text{O} \rightarrow 3\text{ClO}_2^- + 2\text{MnO}_2 + 2\text{OH}^-$

The reducing agent is:

- A) H_2O B) ClO_2^- C) MnO_2 D) MnO_4^-

9. A strong oxidizing agent has a:

- A) weak attraction for electrons B) strong attraction for electrons
C) weak ability to become reduced D) strong ability to become oxidized

10. Consider the following reaction: $3\text{As}_2\text{O}_3 + 4\text{NO}_3^- + 7\text{H}_2\text{O} \rightarrow 6\text{H}_2\text{AsO}_4^- + 4\text{HNO}_3$

The oxidizing agent is:

- A) H^+ B) H_2O C) NO_3^- D) As_2O_3

11. What is a typical characteristic of a strong oxidizing agent?

- A) It is readily oxidized B) It easily loses electrons.
C) It has a negative oxidation number. D) It has a positive reduction potential.

12. Two half-cells are connected via a salt bridge to make a voltaic cell. Which statement about this cell is correct?

- A) Oxidation occurs at the positive electrode (cathode).
B) It is also known as an electrolytic cell.
C) Ions flow through the salt bridge.
D) It requires a power supply to operate.

13. A solution containing an unknown cation reacts spontaneously with both zinc and copper. The unknown cation is:

- A) 1.0M H^+ B) 1.0M Ag^+ C) 1.0M Sr^{2+} D) 1.0M Ni^{2+}

14. As a metal corrodes:

- A) it gains electrons
B) it becomes reduced
C) it acts as a reducing agent
D) its oxidation number decreases

15. As a standard Zn / Ag electrochemical cell operates, in which direction do anions

move and how does the mass of the cathode change?

	Anion Direction	Mass of Cathode
A)	Towards Zn electrode	Increases
B)	Towards Ag electrode	Increases

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