

Ions in Aqueous Solutions and Colligative Properties

practice:

practice:

and the dissolution of each of the following in water, and then determine the

are the equation of each ion produced as well as the total number of moles. and the coast of each ion produced as well as the total number of moles of ions produced. e l'and ammontant chloride

practice:

Will a precipitate form if solutions of potassium sulfate and barium estrate are combined?

As not sonic equation for the reaction. I were the net sonic equation for the reaction.

1 will a precipitate form if solutions of potassium nitrate and magnesium sulfate are ordered! If so, write the net ionic equation for the reaction.

3. Will a precipitate form if solutions of barium chloride and sodium sulfate are combined? If as dentify the spectator ions and write the net ionic equation.

4 water the not some equation for the precipitation of nickel(II) sulfide.

SECTION REVIEW

f, was the equation for the dissolution of 0.5 mol Se(NOs)s in water. How many moles of greener ices and nitrate ions are produced?

as not of Sr*2 ions and 1 mol of NOs* ions are produced.

1 Will a procipitate form if solutions of magnesium acetate and strontium chloride are

No. The salts that could potentially form, magnesium chloride and strontium acetate. - Decision II are so builded.

), what determines whether a molecular compound will be ionized in a polar solvent? 108 / If the attraction of polar solvent molecules is stronger than a covalent bond in the solute, me molecule will break into sons.

4 Explain why HCl is a strong electrolyte and HF is a weak electrolyte.

ANSI HCl is essentially 100% ionized in water solution, whereas HF is only slightly ionized because its H-F covalent bond is strong.

5 for each of the following pairs, tell which solution contains the larger total modestration of foms.

a 0.10 M HCl and 0.05 M HCl

8: 0.10 M HCl and 0.10 M HF

= 0.10 M HCl and 0.10 M CaCl₂

ANS I at 0.10 M HCl; HCl is a strong acid, so the HCl in solution is completely converted to 8:0'(sq) and Cl (sq). The HCl solution with the higher concentration would have more ions. k.4.10 M HCl; HF is a weak acid, so only some of the HF molecules ionize.

c.0.10 M CaCli; In 0.10 M solution, HCl (a strong acid) forms 0.10 M H₂O' and 0.10 M Cl⁻ to are a total ion concentration of 0.20 M. CaCly is a soluble ionic salt that forms 0.10 M Ca²² and 0.20 M ClT to give a total ion concentration of 0.30 M.

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Melvin D. Joesten, James L. Wood

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