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## pH and pOH

The pH of a solution indicates how acidic or basic that solution is.

pH range 0 to 7 acidic

7 neutral

7-14 basic

Since  $[H^+][OH^-] = 10^{-14}$  at 25 °C, if  $[H^+]$  is known, the  $[OH^-]$  can be calculated and vice versa.

$$pH = -\log [H^+]$$

$$pOH = -\log [OH^-]$$

$$\text{So if } [H^+] = 10^{-9} \text{ M, } pH = 9.$$

$$\text{So if } [OH^-] = 10^{-5} \text{ M, } pOH = 5.$$

Together,  $pH + pOH = 14$ .

Complete the following chart.

|     | $[H^+]$              | pH | $[OH^-]$             | pOH | Acidic or Basic |
|-----|----------------------|----|----------------------|-----|-----------------|
| 1.  | $10^{-1} \text{ M}$  | 1  | $10^{-13} \text{ M}$ | 13  | Acidic          |
| 2.  |                      | 2  |                      |     |                 |
| 3.  |                      |    | $10^{-6} \text{ M}$  |     |                 |
| 4.  | $10^{-8} \text{ M}$  |    |                      |     |                 |
| 5.  |                      |    |                      | 11  |                 |
| 6.  |                      | 12 |                      |     |                 |
| 7.  |                      |    | $10^{-7} \text{ M}$  |     |                 |
| 8.  | $10^{-11} \text{ M}$ |    |                      |     |                 |
| 9.  |                      |    |                      | 13  |                 |
| 10. |                      | 8  |                      |     |                 |

# South Pasadena Ph And Poh Answers

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