

SUZUKI LT-Z400K3 (MODELO 2003)



LADO DERECHO



LADO IZQUIERDO

* Puede haber diferencias entre las fotografías y el vehículo real dependiendo de los mercados.

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El número de serie del bastidor o V.I.N. (Número de Identificación del Vehículo) ① está estampado en el lado izquierdo del tubo del bastidor. El número de serie del motor ② está situado en el lado derecho del cárter. Estos números son especialmente necesarios para registrar la máquina y pedir recambios.



RECOMENDACIÓN DE COMBUSTIBLE Y ACEITE COMBUSTIBLE (Para CANADÁ y EE.UU.)

- Utilice solamente gasolina sin plomo de 87 octanos como mínimo ($R/2 + M/2$) o de 91 octanos o más según el Método de Investigación.
- SUZUKI recomienda a los clientes utilizar gasolina sin plomo y sin alcohol siempre que sea posible.
- Se puede utilizar gasolina que contenga MTBE (Methyl Tertiary Butyl Ether).
- Se puede utilizar mezcla de gasolina/alcohol, siempre que ésta no contenga más del 10% de etanol. El combustible de gasolina/alcohol puede contener hasta 5% de metanol si tiene también los codisolventes e inhibidores de corrosión apropiados.
- Si el rendimiento del vehículo no es satisfactorio cuando utiliza la mezcla de gasolina/alcohol, deberá cambiar a gasolina sin plomo ni alcohol.
- Si no sigue estos consejos podría anular la garantía del vehículo. Consulte con su proveedor de combustible y asegúrese de que el combustible que **piensa utilizar** cumple los requisitos indicados más arriba.

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web limiting reagent and percent yield practice worksheet answer key limiting reactant yield practice worksheet 1 methanol CH_3OH can be produced by the following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$ calculate the theoretical yield of CH_3OH if 68.5 g of CO is reacted with 8.6 g of H_2

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web limiting reactant and percent yield practice name 1 consider the following reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{}_3\text{PO}_4 + \text{N}_2\text{O}$ which reactant is limiting assuming we started with 30.0 grams of ammonium nitrate and 50.0 grams of sodium phosphate what is the

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web limiting reagents and percentage yield worksheet 1 consider the reaction $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{CO}_2 + \text{I}_2$ a 80.0 grams of iodine v oxide I_2O_5 reacts with 28.0 grams of carbon monoxide CO determine the mass of iodine I_2 which could be produced 80 g I_2O_5 1 mol I_2O_5 1 mol I_2 xs 1 333.8 g I_2O_5 1 mol I_2O_5 28 g CO 1 mol CO

7 3 limiting reactant and percent yield problems - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3