Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is
A) the reactant of a chemical process that is not consumed completely
B) the product of a chemical process that is consumed completely
C) the product of a chemical process that is not consumed completely
D) the reactant of a chemical process that is consumed completely
E) none of the above

ANS: D

1. Consider the following reaction: $Ca(OH)_{2(aq)} + Al_2(SO_4)_{3(aq)} \rightarrow CaSO_{4(s)} + Al(OH)_{3(s)}$
When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is
A) 1:1 B) 2:1 C) 2:3 D) 1:2 E) 3:1

ANS: E

1. Sodium phosphate, Na_3PO_4, can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation.
$3NaOH + H_3PO_4 \rightarrow Na_3PO_4 + 3H_2O$
If 2.5 mol of H_3PO_4 are reacted, the number of moles of water that will be produced is
A) 1.0 B) 2.5 C) 3.0 D) 5.0 E) 7.5

ANS: E

1. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process.
$2KClO_3 \rightarrow 2KCl + 3O_2$
The number of moles of $KClO_3$ required to produce 5.0 mol of oxygen is
A) 1.0 B) 2.0 C) 3.0 D) 3.3 E) 5.0

ANS: D

1. Determine the number of grams of CO_2 that can be produced from 48.0 g of the hydrocarbon C_3H_8, according to the following equation.
$2C_3H_8 + 9O_2 \rightarrow 6CO_2 + 8H_2O$
A) 37.6 g B) 50.2 g C) 87.9 g D) 151 g E) 198.4 g

1. The limiting reactant in a chemical reaction is the _____ that is _____.
A) Reactant, in excess B) Product, left over

Solution Stoichiometry Multiple Choice Answer

Stephen DeMeo

Solution Stoichiometry Multiple Choice Answer:

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