

Chapter 1 • Introduction

P1.1 A gas at 20°C may be *rarefied* if it contains less than 10^{12} molecules per mm^3 . If Avogadro's number is $6.023\text{E}23$ molecules per mole, what air pressure does this represent?

Solution: The mass of one molecule of air may be computed as

$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.023\text{E}23 \text{ molecules/g} \cdot \text{mol}} = 4.81\text{E}-23 \text{ g}$$

Then the density of air containing 10^{12} molecules per mm^3 is, in SI units,

$$\begin{aligned} \rho &= \left(10^{12} \frac{\text{molecules}}{\text{mm}^3} \right) \left(4.81\text{E}-23 \frac{\text{g}}{\text{molecule}} \right) \\ &= 4.81\text{E}-11 \frac{\text{g}}{\text{mm}^3} = 4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \end{aligned}$$

Finally, from the perfect gas law, Eq. (1.13), at $20^\circ\text{C} = 293 \text{ K}$, we obtain the pressure:

$$p = \rho RT = \left(4.81\text{E}-5 \frac{\text{kg}}{\text{m}^3} \right) \left(287 \frac{\text{m}^2}{\text{s}^2 \cdot \text{K}} \right) (293 \text{ K}) = \mathbf{4.0 \text{ Pa}} \quad \text{Ans.}$$

Solutions Manual Elger Fluid Mechanics 9th Edition

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