

Chemical Equilibrium

Multiple Choice Practice

1. Chemical reactions are in a state of dynamic equilibrium only when
- the rate of the forward reaction equals that of the backward reaction.
 - the concentration of reactants and products are equal
 - the activation energies of the forward and backward reactions are equal
 - the reaction involves no enthalpy change.
2. Under the conditions used industrially, ethene and steam react as follows:



Which set of conditions would give the best yield of ethanol at equilibrium?

- high temperature, low pressure
 - high temperature, high pressure
 - low temperature, high pressure
 - low temperature, low pressure.
3. $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightleftharpoons 2\text{NH}_3(\text{g})$
- In the Haber process, illustrated by the equation, the reaction never reaches equilibrium because
- the ammonia is constantly removed
 - an iron catalyst is used
 - high pressure is used
 - the temperature is held at above 500°C .

4. The equation refers to the preparation of methanol from synthesis gas.



The formation of methanol is favoured by

- high pressure and low temperature
 - high pressure and high temperature
 - low pressure and low temperature
 - low pressure and high temperature.
5. $\text{ICl}(\text{l}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{ICl}_3(\text{s})$
- $$\Delta H = -46 \text{ kJ mol}^{-1}$$
- Which of the following sets of changes will cause the greatest increase in the proportion of solid in the above equilibrium?
- | | Temperature | Pressure |
|---|-------------|----------|
| A | decrease | decrease |
| B | decrease | increase |
| C | increase | decrease |
| D | increase | increase |
6. When a reversible chemical reaction is at equilibrium,
- the concentration of reactants and products remain equal
 - the forward reaction is unable to continue
 - the concentration of reactants and products remain constant
 - the forward and reverse reactions proceed at different rates.

Static Equilibrium Multiple Choice

Ying-Ying Zheng



Static Equilibrium Multiple Choice:

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