

Name Key Period _____

Worksheet: Calculating Empirical & Molecular Formulas

1. The empirical formula for the compound having the formula $H_4C_2O_4$ is:

- [A] C_2H_2 [B] CO_2H [C] COH [D] $C_2O_4H_4$ [E] COH_2

2. Calculate the empirical formula of a compound that is 85.6% C and 14.4% H (by mass).

- [A] CH_2 [B] CH [C] C_2H_4 [D] C_2H_6 [E] C_3H_8

$$C: \frac{85.6}{12.01} = 7.12 \quad H: \frac{14.4}{1.01} = 14.25 = \frac{14.25}{7.12} = 2$$

3. A compound is analyzed and found to contain 12.1% carbon, 56.2% oxygen, and 71.7% chlorine (by mass). Calculate the empirical formula of this compound.

- [A] $COCl_2$ [B] CO_2Cl_2 [C] CO_2Cl [D] $COCH$ [E] $COCl$

$$C: \frac{12.1}{12.01} = 1 \quad O: \frac{56.2}{16.00} = 3.51 \quad Cl: \frac{71.7}{35.45} = 2$$

4. A compound contains 40.0% carbon, 6.7% hydrogen, and 53.3% oxygen (by mass). Calculate the empirical formula.

- [A] C_2H_2O [B] CH_2O [C] CH_4O [D] $C_2H_4O_2$ [E] $C_2H_6O_2$

$$C: \frac{40.0}{12.01} = 3.33 \quad H: \frac{6.7}{1.01} = 6.63 \quad O: \frac{53.3}{16.00} = 3.33 = \frac{6.63}{3.33} = 2$$

5. A compound contains 25.94% N and 74.06% O (by mass). What is the empirical formula?

$$\frac{25.94g N}{14.01} = \frac{1.85}{1.84} = 1 \times 2 = 2$$

$$\frac{74.06g O}{16.00} = \frac{4.63}{1.84} = 2.5 \times 2 = 5$$



6. Determine the empirical formula of a compound containing 54.2% F and 45.8% S (by mass).

$$\frac{54.2g F}{19.00g/mol} = \frac{2.85}{1.43} mol F = 1.99 \sim 2$$

$$\frac{45.8g S}{32.07g/mol} = \frac{1.43}{1.43} mol S = 1$$



7. A compound has 40.68% carbon, 5.12% hydrogen, and 54.20% oxygen (by mass). Calculate its empirical formula.

$$\frac{40.68g C}{12.01g/mol} = \frac{3.387}{3.387} \quad \frac{5.12g H}{1.01g/mol} = \frac{5.069}{3.387} \quad \frac{54.20g O}{16.00g/mol} = \frac{3.387}{3.387}$$

$$= 1 \times 2 = 2 \quad = 1.496 \times 2 = 3 \quad = 1 \times 2 = 2$$



8. A 7.33-g sample of lanthanum, La, combines with oxygen to give 10.29 g of the oxide. Calculate the empirical formula of this oxide.

$$10.29g - 7.33g La = 2.96g O$$

$$\frac{7.33g La}{138.91g/mol} = .0527 \quad \frac{2.96g O}{16.00g/mol} = .1850 = \frac{.1850}{.0527} = 3.5 \times 2 = 7$$

$$= 1 \times 2 = 2$$



9. Calculate the molecular formula of a compound with the empirical formula CH_2O and a molar mass of 150 g/mol.

- [A] $C_7H_6O_3$ [B] $C_5H_{10}O_5$ [C] $C_3H_6O_3$ [D] $C_4H_8O_4$ [E] $C_6H_{12}O_6$

$$30 \quad 150 \div 30 = 5$$

Skill Practice 3moles And Formulas Answer Key

**Donald J. Wink, Sharon Fetzer-
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Table of Contents Skill Practice 3moles And Formulas Answer Key

1. Understanding the eBook Skill Practice 3moles And Formulas Answer Key
 - The Rise of Digital Reading Skill Practice 3moles And Formulas Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Skill Practice 3moles And Formulas Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Skill Practice 3moles And Formulas Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Skill Practice 3moles And Formulas Answer Key
 - Personalized Recommendations
 - Skill Practice 3moles And Formulas Answer Key User Reviews and Ratings

- Skill Practice 3moles And Formulas Answer Key and Bestseller Lists
- 5. Accessing Skill Practice 3moles And Formulas Answer Key Free and Paid eBooks
 - Skill Practice 3moles And Formulas Answer Key Public Domain eBooks
 - Skill Practice 3moles And Formulas Answer Key eBook Subscription Services
 - Skill Practice 3moles And Formulas Answer Key Budget-Friendly Options
- 6. Navigating Skill Practice 3moles And Formulas Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Skill Practice 3moles And Formulas Answer Key Compatibility with Devices
 - Skill Practice 3moles And Formulas Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Skill Practice 3moles And Formulas Answer Key
 - Highlighting and Note-Taking Skill Practice 3moles And Formulas Answer Key
 - Interactive Elements Skill Practice 3moles And Formulas Answer Key
- 8. Staying Engaged with Skill Practice 3moles And Formulas Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Skill Practice 3moles And Formulas Answer Key
- 9. Balancing eBooks and Physical Books Skill Practice 3moles And Formulas Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Skill Practice 3moles And Formulas Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Skill Practice 3moles And Formulas Answer Key
 - Setting Reading Goals Skill Practice 3moles And Formulas Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Skill Practice 3moles And Formulas Answer Key
 - Fact-Checking eBook Content of Skill Practice 3moles And Formulas Answer Key
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web when responding to this competency based interview question you would need to provide a specific example of where you have already demonstrated the core skills necessary surrounding the competency of teamwork what are the likely competencies to be assessed during my united nations interview

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web unlike most job interviews the un hiring panel relies on competency based questions given that this organization incredibly takes pride in the quality of its staff most of the questions that we will cover in this article will therefore be behavioral or criterion based seeking to uncover some of your past behaviors and experiences under

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web questions should be asked exactly as stated in the ecdi2030 with no changes to wording except for one question that follows a skip rule as explained later in this manual interviewers must ask all 20 questions the fact that some questions might sound too hard or too easy can generate reactions from respondents

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web explore the resources below to prepare yourself for the competency based interviewing style used in the un system guidance read these tips on preparing for competency based interviews