Moles and calculations

Define moles: It is the amount of substance which is equal to the Avogadro's number of particles.

Avogadro's number or Avogadro's Constant = 6.02 x 10²⁰ particles Formulas for calculation of moles:

1. mole = mass

Mass of one mole

Mass of one mole = molar mass = Mr or Ar

2. mole of any gas = Volume of gas in cm³ Volume of gas in dm³

Volume of one mole of gas Volume of one mole of gas

Volume of one mole of any gas = molar gas volume = 24000 cm3 or 24 dm3

mole of any substance = volume in dm³ x concentration in mol dm³ (In the form of solution)

Volume in dm³ = Volume of solution in cm³

4. percentage yield = Experimental yield × 100

5. percentage purity = Mass of pure substance × 100

7. molecular formula = n x empirical formula

Where n = Empirical formula mass

Skill Practice 40 Moles And Formula Practice

Albert A Gayle

Skill Practice 40 Moles And Formula Practice:

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equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands equilibrium calculations with complexes and applications of complexation chemistry Oxidation reduction equilibria including equilibrium calculations graphical approaches and applications Gas liquid and solid liquid equilibrium with expanded coverage of the effects of global climate change Other topics including chemical kinetics of aquatic systems surface chemistry and integrative case studies For advanced senior undergraduates and first year graduate students in environmental engineering courses A Problem Solving Approach to Aquatic Chemistry serves as an invaluable learning resource on the topic with a variety of helpful learning elements included throughout to ensure information retention and the ability to apply covered concepts in practical settings

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