

Name \_\_\_\_\_

Chemistry Period \_\_\_\_\_

### Worksheet Specific Heat

Specific Heat of water =  $4.18 \text{ J/g}^\circ\text{C}$

1. A sample of mercury is heated from  $25.5^\circ\text{C}$  to  $52.5^\circ\text{C}$ . In the process, 3050J of heat are absorbed. What mass of mercury was in the sample? The specific heat of mercury is  $0.140 \text{ J/g}^\circ\text{C}$ . **ANS: 807 g**
2. A block of aluminum with a mass of 140g is cooled from  $98.4^\circ\text{C}$  to  $62.2^\circ\text{C}$  with a release of 1137J of heat. From these data, calculate the specific heat of aluminum. **ANS: 0.224**
3. A total of 226 J of heat are absorbed as 58.3 g of lead is heated from  $12.0^\circ\text{C}$  to  $42.0^\circ\text{C}$ . From this data, what is the specific heat of lead? **ANS: 0.129**
4. What is the change in heat energy when 132g of copper at  $32.2^\circ\text{C}$  is raised to  $45.0^\circ\text{C}$ ? The specific heat of copper is  $0.385 \text{ J/g}^\circ\text{C}$ . **ANS: 650 J**
5. What is the change in heat energy when 40.0g of water at  $25.0^\circ\text{C}$  is raised to  $60.0^\circ\text{C}$ ? **ANS: 5852 J**
6. What is the change in heat energy when 75.0 of water at  $9.8^\circ\text{C}$  is raised to  $22.4^\circ\text{C}$ ? **3950 J**
7. What is the change in heat energy when 125g of water at  $10.0^\circ\text{C}$  is raised to  $90.0^\circ\text{C}$ ? **ANS: 41.8 kJ**
8. What is the change in heat energy when 64.82 g of aluminum metal at  $100.0^\circ\text{C}$  is cooled to  $82.0^\circ\text{C}$ ? The specific heat of aluminum is  $0.897 \text{ J/g}^\circ\text{C}$ . **ANS: 1047**
9. What is the change in temperature if 24g of water at  $50^\circ\text{C}$  absorbs 985 J of heat energy? **ANS: 9.82**
10. What is the mass of a sample of iron if that sample lost 2300J of heat energy when it cooled from  $80^\circ\text{C}$  to  $30^\circ\text{C}$ ? The specific heat of iron is  $0.449 \text{ J/g}^\circ\text{C}$ . **ANS: 102 g**
11. What is the specific heat of metal if its mass is 26.86g and it requires 418.6 J of heat energy to raise its temperature from  $27.4^\circ\text{C}$  to  $67.3^\circ\text{C}$ ? **ANS: 0.391**
12. How much heat energy is needed to raise the temperature of 59.7g of cadmium from  $25^\circ\text{C}$  to  $100^\circ\text{C}$ ? The specific heat of cadmium is  $0.231 \text{ J/g}^\circ\text{C}$ . **ANS: 1034 J**

# Skills Math Skills Specific Heat Answers

**Rod Powers, Jennifer Lawler**



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