

CHAPTER 2 WORKSHEET: CHEMISTRY OF LIFE

1. Chapter 2, Section 1: Nature of Matter

- a. Matter: anything that occupies space. Ex: Desk, Human, Air, ect.
- b. Atom: Smallest part of matter. It **can not** be broken down into smaller pieces. Composed of 3 particles: Protons (+), Electrons (-), and Neutrons (n)
- c. Element: Substance that is only made of one kind of atom. Ex: Hydrogen, Oxygen, ect.
 - i. C, H, N, O, P, S
- d. Compound: Two or more different atoms or elements joined together. Ex: Salt =
- e. Covalent Bond: Two or more atoms sharing electrons to form a molecule. Covalent bonds make molecules, ex: Water, Carbon Dioxide, and Oxygen Gas
- f. Hydrogen Bond: A weak chemical attraction between **polar** molecules. Ex: Two water molecules
- g. Ionic Bond: Gain or Loss of electrons that create a molecule
- h. Ion: An atom that has gained (-) electrons or (loss) electrons (+)

2. Chapter 2, Section 2: Water and Solutions

- a. How does water help maintain homeostasis?
 - i. Heats slowly
 - ii. Retains heat
 - iii. Carries Heat away through water vapor (sweating)
 - iv. Universal solvent (anything except for a polar substance can dissolve in it)
- b. What is the difference between adhesion & cohesion?
 - i. Adhesion: different substances
 - ii. Cohesion: alike substances
- c. What is the pH range for an Acid? Base?
 - i. Acid = 0 – 6.9
 - ii. Neutral = 7
 - iii. Base = 7.1 – 14
- d. What type of substance is an example of an acid? Of a base?
 - i. Acids: Lemon, Vinegar, Milk
 - ii. Bases: Household cleaning material

3. Chapter 2, Section 3: Chemistry of Cells

- a. Carbohydrates: Organic substances, made up of carbon, hydrogen and oxygen atoms. Always in the ratio of 1:2:1. Key source of energy.
 - i. They are made up of Monosaccharide: Single sugars (glucose)
 - ii. Disaccharides are double sugars: Sucrose
 - iii. Polysaccharides: Starch and cellulose

The Chemistry Of Life Chapter 2 Answers

Janet Rae-Dupree, Pat DuPree



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