

TITRATION PRACTICE WORKSHEET

- If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl? (0.04M)
- If it takes 25 mL of 0.05M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution? ($4 \times 10^{-3} \text{ M}$)
- If it takes 50 mL of 0.5 M KOH solution to completely neutralize 125 mL of sulphuric acid solution (H_2SO_4), what is the concentration of the acid solution? (0.1 M)
- Can I titrate a solution of unknown concentration with another solution of unknown concentration and still get a meaningful answer? Explain your answer in a few sentences.
- Explain the difference between an endpoint and equivalence point in a titration.
- It takes 26.23 mL of a 1.008 M NaOH solution to neutralize a solution of 5g of an unknown monoprotic acid in 150.2 mL of solution. What is the molecular weight of the unknown? (192.3 g/mol)
- A solution is made by dissolving 12.50 g of NaOH in water to produce 2.0 L of solution. What is the pH of this solution? Hint: $\text{pH} = 14 - \text{pOH}$ and $\text{pOH} = -\log [\text{OH}^-]$
- 17.85 mL of nitric acid neutralizes 25.00 mL of 0.150 mol/L NaOH. What is the concentration of the nitric acid? (0.210 mol/L)
- What volume of 1.015 mol/L magnesium hydroxide is needed to neutralize 40.0 mL of 1.60 mol/L hydrochloric acid? (31.5 mL)
- What volume of 0.150 M HCl is needed to neutralize each solution listed below?
 - 25.0 mL of 0.135 M NaOH (22.5 mL)
 - 20.0 mL of 0.185 M NH_4OH (24.7 mL)
 - 80.0 mL of 0.0045 M $\text{Ca}(\text{OH})_2$ (4.8 mL)
- What concentration of sodium hydroxide is required for each neutralization reaction?
 - 37.82 mL of NaOH neutralizes 15.00 mL of 0.250 M HF (0.0991 M)
 - 21.56 mL of NaOH neutralizes 20.00 mL of 0.145 M H_2SO_4 (0.269 M)
 - 14.27 mL of NaOH neutralizes 25.00 mL of 0.105 M H_3PO_4 (0.552 M)
- A 25.0 mL sample of sulphuric acid is completely neutralized by adding 32.8 mL of 0.116 M ammonia solution. Ammonium sulphate is formed. What is the concentration of the sulphuric acid? ($7.6 \times 10^{-2} \text{ M}$)
- The following data were collected during a titration. Calculate the concentration of the sodium hydroxide solution. (0.107 M)

Titration Data	
Volume of HCl	10.00 mL
Final volume of NaOH	23.08 mL
Initial volume of NaOH	1.06 mL
Concentration of HCl	0.235 M
- In a titration experiment, 25.0 mL of an aqueous solution of sodium hydroxide was required to neutralize 50.0 mL of a 0.010 M hydrochloric acid. What is the concentration of the sodium hydroxide solution? (0.020 M)

Titration Problem And Solution Practice Answers

Sebastian Brünink



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Volumetric Analysis Ralf Martens-Menzel, Lena Harwardt, Hanns-Jürgen Krauss, 2025-05-19 This standard work on volumetric analysis based on the 20th German edition provides comprehensive information on the theory of acid base titration redox titration complexation titration and precipitation titration with both classical and instrumental indication of the equivalence point Many applications are described and explained in detail with examples in pharmaceutical and environmental analysis

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