

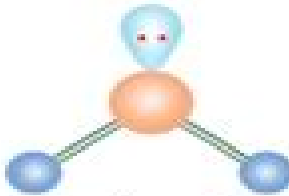
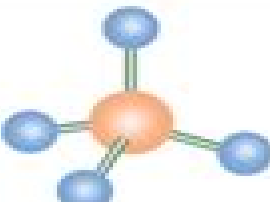
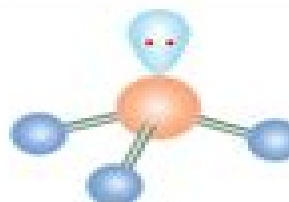
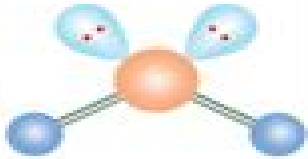

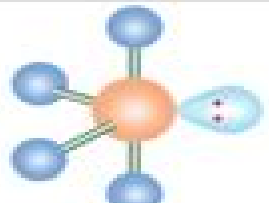



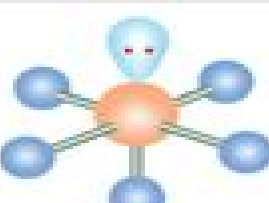
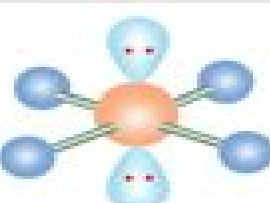

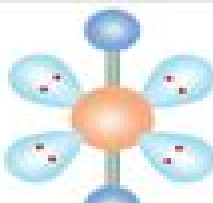


VSEPR Theory Chart

Number of Electron Groups	Lone Pairs = 0	Lone Pairs = 1	Lone Pairs = 2	Lone Pairs = 3	Lone Pairs = 4
2	 <p>Linear</p>				
3	 <p>Trigonal Planar</p>	 <p>Angular or Bent</p>			
4	 <p>Tetrahedral</p>	 <p>Trigonal Pyramidal</p>	 <p>Angular or Bent</p>		
5	 <p>Trigonal Bipyramidal</p>	 <p>Seesaw</p>	 <p>T-shaped</p>	 <p>Linear</p>	
6	 <p>Octahedral</p>	 <p>Square Pyramidal</p>	 <p>Square Planar</p>	 <p>T-shaped</p>	 <p>Linear</p>

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