

Physics

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(Chapter 12) (Kinetic Theory) (Exercises)
(Class 11)

Question 12.1:

Estimate the fraction of molecular volume to the actual volume occupied by oxygen gas at STP. Take the diameter of an oxygen molecule to be $3 \text{ \AA}$.

Answer 12.1:

Diameter of an oxygen molecule, $d = 3 \text{ \AA}$

Radius, $r = \frac{d}{2} = \frac{3}{2} \text{ \AA} = 1.5 \text{ \AA} = 1.5 \times 10^{-8} \text{ cm}$

Actual volume occupied by 1 mole of oxygen gas at STP = 22400 cm^3

Molecular volume of oxygen gas, $V = \frac{4}{3} \pi r^3 N$

Where, N is Avogadro's number = 6.023×10^{23} molecules/mole

$$\therefore V = \frac{4}{3} \times 3.14 \times (1.5 \times 10^{-8})^3 \times 6.023 \times 10^{23} = 8.51 \text{ cm}^3$$

$$\text{Ratio of the molecular volume to the actual volume of oxygen} = \frac{8.51}{22400} = 3.8 \times 10^{-4}$$

Question 12.2:

Molar volume is the volume occupied by 1 mol of any (ideal) gas at standard temperature and pressure (STP: 1 atmospheric pressure, 0°C). Show that it is 22.4 litres.

Answer 12.2:

The ideal gas equation relating pressure (P), volume (V), and absolute temperature (T) is given as:

$$PV = nRT$$

Where,

R is the universal gas constant = $8.314 \text{ J mol}^{-1} \text{ K}^{-1}$

n = Number of moles = 1

T = Standard temperature = 273 K

P = Standard pressure = $1 \text{ atm} = 1.013 \times 10^5 \text{ Nm}^{-2}$

$$\therefore V = \frac{nRT}{P} = \frac{1 \times 8.314 \times 273}{1.013 \times 10^5}$$

$$= 0.0224 \text{ m}^3 = 22.4 \text{ litres}$$

Hence, the molar volume of a gas at STP is 22.4 litres.

Question 12.3:

Figure 12.8 shows plot of PV/T versus P for $1.00 \times 10^{-3} \text{ kg}$ of oxygen gas at two different temperatures.

(a) What does the dotted plot signify?

(b) Which is true: $T_1 > T_2$ or $T_1 < T_2$?

(c) What is the value of PV/T where the curves meet on the y-axis?

(d) If we obtained similar plots for $1.00 \times 10^{-3} \text{ kg}$ of hydrogen, would we get the same value of PV/T at the point where the curves meet on the y-axis? If not, what mass of hydrogen yields the same value of PV/T (for low pressure high temperature region of the plot)?

(Molecular mass of $\text{H}_2 = 2.02 \text{ u}$, of $\text{O}_2 = 32.0 \text{ u}$, $R = 8.31 \text{ J mol}^{-1} \text{ K}^{-1}$.)



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