

MULTIPLE CHOICE. Choose the one alternative that best completes the statement or answers the question.

- 1) Consider the following reaction:



The average rate of appearance of B is given by $\Delta[B]/\Delta t$. Comparing the rate of appearance of B and the rate of disappearance of A, we get $\Delta[B]/\Delta t = \underline{\hspace{1cm}} \times (-\Delta[A]/\Delta t)$.

- A) $-2/3$ B) $-3/2$ C) $+2/3$ D) $+3/2$ E) $+1$
- 2) A burning splint will burn more vigorously in pure oxygen than in air because
- A) nitrogen is a reactant in combustion and its low concentration in pure oxygen catalyzes the combustion.
 B) oxygen is a reactant in combustion and the concentration of oxygen is higher in pure oxygen than it is in air.
 C) oxygen is a product of combustion.
 D) nitrogen is a product of combustion and the system reaches equilibrium at a lower temperature.
 E) oxygen is a catalyst for combustion.
- 3) Of the following, all are valid units for a reaction rate except _____.
- A) mol/L B) M/s C) mol/hr D) mol/L-hr E) g/s

A flask is charged with 0.124 mol of A and allowed to react to form B according to the reaction $A(g) \rightarrow B(g)$. The following data are obtained for [A] as the reaction proceeds:

| Time (s) | 1 | 10 | 20 | 30 | 40 |
|------------|-------|-------|-------|-------|-------|
| Moles of A | 0.124 | 0.110 | 0.088 | 0.073 | 0.054 |

- 4) The average rate of disappearance of A between 10 s and 20 s is _____ mol/s.
- A) 1.1×10^{-3} B) 4.4×10^{-3} C) 454 D) 9.90×10^{-3} E) 2.2×10^{-3}

The peroxydisulfate ion ($S_2O_8^{2-}$) reacts with the iodide ion in aqueous solution via the reaction:



An aqueous solution containing 0.050 M of $S_2O_8^{2-}$ ion and 0.072 M of I^- is prepared, and the progress of the reaction followed by measuring $[I^-]$. The data obtained is given in the table below.

| Time (s) | 0 | 400 | 800 | 1200 | 1600 |
|-------------|-------|-------|-------|-------|-------|
| $[I^-]$ (M) | 0.072 | 0.057 | 0.046 | 0.037 | 0.029 |

- 5) The concentration of $S_2O_8^{2-}$ remaining at 800 s is _____ M.
- A) 0.015 B) 0.041 C) 0.076 D) 4.00×10^{-3} E) 0.046
- 6) Of the units below, _____ are appropriate for a first-order reaction rate constant.
- A) $L \cdot mol^{-1} s^{-1}$ B) s^{-1} C) $M^{-1} s^{-1}$ D) mol/L E) $M s^{-1}$

Timberlake Chemistry Practice Test Ch 12

Didier Musso



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Table of Contents Timberlake Chemistry Practice Test Ch 12

1. Understanding the eBook Timberlake Chemistry Practice Test Ch 12
 - The Rise of Digital Reading Timberlake Chemistry Practice Test Ch 12
 - Advantages of eBooks Over Traditional Books
2. Identifying Timberlake Chemistry Practice Test Ch 12
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Timberlake Chemistry Practice Test Ch 12
 - User-Friendly Interface
4. Exploring eBook Recommendations from Timberlake Chemistry Practice Test Ch 12
 - Personalized Recommendations
 - Timberlake Chemistry Practice Test Ch 12 User Reviews and Ratings
 - Timberlake Chemistry Practice Test Ch 12 and Bestseller Lists
5. Accessing Timberlake Chemistry Practice Test Ch 12 Free and Paid eBooks
 - Timberlake Chemistry Practice Test Ch 12 Public Domain eBooks
 - Timberlake Chemistry Practice Test Ch 12 eBook Subscription Services
 - Timberlake Chemistry Practice Test Ch 12 Budget-Friendly Options

6. Navigating Timberlake Chemistry Practice Test Ch 12 eBook Formats
 - ePub, PDF, MOBI, and More
 - Timberlake Chemistry Practice Test Ch 12 Compatibility with Devices
 - Timberlake Chemistry Practice Test Ch 12 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Timberlake Chemistry Practice Test Ch 12
 - Highlighting and Note-Taking Timberlake Chemistry Practice Test Ch 12
 - Interactive Elements Timberlake Chemistry Practice Test Ch 12
8. Staying Engaged with Timberlake Chemistry Practice Test Ch 12
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Timberlake Chemistry Practice Test Ch 12
9. Balancing eBooks and Physical Books Timberlake Chemistry Practice Test Ch 12
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Timberlake Chemistry Practice Test Ch 12
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Timberlake Chemistry Practice Test Ch 12
 - Setting Reading Goals Timberlake Chemistry Practice Test Ch 12
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Timberlake Chemistry Practice Test Ch 12
 - Fact-Checking eBook Content of Timberlake Chemistry Practice Test Ch 12
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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