Activity B:	Get the Gizmo ready:	5.3
Determining concentration	 Click Reset. Select 1.00 M NaOH for the Burette and Mystery H₂SO₄ for the Flask. Select Bromthymol blue for the Indicator. 	

Introduction: Adding a drop of strong acid or base into a neutralized solution is similar to adding a drop of strong acid or base to water—it causes an abrupt change in pH. By using an appropriate indicator, a chemist can tell when a solution is neutralized by monitoring its color.

Question: How is titration used to determine an unknown concentration?

Measure: Titrate the sulfuric acid analyte (H₂SO₄) with the sodium hydroxide titrant (NaOH).

How much 1.00 M NaOH is needed to neutralize the H₂SO₄ solution?

21.6 mL

Interpret: The balanced equation for the reaction of H₂SO₄ and NaOH is given at bottom right.

Based on this equation, how many moles of NaOH react with 1 mole of H₂SO₄?

2

- Manipulate: Recall that molarity is equal to the number of moles of a substance dissolved in one liter of solution: molarity = moles + volume.
 - A. Write an equation for determining the number of moles of NaOH that are added to the flask based on [NaOH] and volume of NaOH titrant (ml. NaOH):

Moles NaOH =
$$[NaOH] * mL NaOH/1,000mL$$

B. Write a similar expression for the number of moles of H₂SO₄ in the flask based on [H₂SO₄] and the volume of H₂SO₄ (mL).

Moles
$$H_2SO_4 = [H_2SO_4] * mL H_2SO_4/1,000 mL$$

C. Because there are twice as many moles of NaOH as moles of H₂SO₄ in this reaction, you can say:

Substitute your expressions from 3A and 3B into this equation and solve for [H₂SO₄]:

Titration Virtual Activity Answer Key

Hussin A.Rothana

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