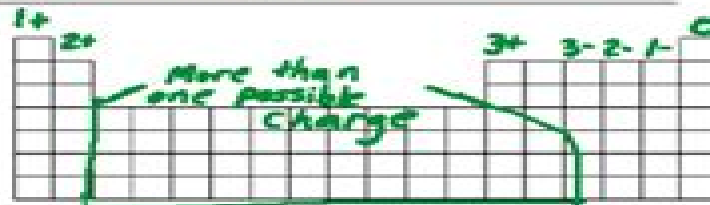


Ionic Bonding Review

Define the following terms.

1. **Electronegativity** *This is the ability of an atom of a given element to attract the electrons in a bond.*
2. **valence electrons** *These are the electrons in the highest energy level of an atom of a given element. These are the electrons which will be shared in a covalent bond or transferred in an ionic bond.*
3. **octet** *This represents the eight valence electrons which the atoms of almost all of the elements have to reach. They will either share electrons or transfer electrons to achieve the octet. The exceptions are the first five elements (hydrogen, helium, lithium, beryllium and boron) which only need two.*
4. **ionic bond** *A bond formed by the attraction between cations (positive ions) and anions (negative ions). Electrons are transferred from the cation to the anion as they form an ionic compound.*
5. Write the charges for the groups on the Periodic Table.
6. Explain why atoms form bonds? *Bonds form so that the system can find the lowest energy state. For ionic compounds this means that the positive and negative ions have been neutralized and the atoms have gained or lost electrons to have the full octet of valence electrons.*



7. Each problem consists of three parts:

1st cation/anion and the number needed to create a neutral compound;

2nd the formulas of the neutral compound, and

3rd the name.

- | | | | | | |
|-----|---------------------------|---|---|---|---------------------------------|
| (A) | <u>2</u> Li^+ | + <u>1</u> O^{2-} | → | <u>Li_2O</u> | <u>lithium oxide</u> |
| (B) | <u>2</u> Fe^{3+} | + <u>3</u> CrO_4^{2-} | → | <u>$\text{Fe}_2(\text{CrO}_4)_3$</u> | iron (III) chromate |
| (C) | <u>1</u> Sr^{2+} | + <u>2</u> NO_3^- | → | <u>$\text{Sr}(\text{NO}_3)_2$</u> | <u>strontium nitrate</u> |
| (D) | <u>3</u> Ca^{2+} | + <u>2</u> N^{3-} | → | <u>Ca_3N_2</u> | calcium nitride |
| (E) | <u>1</u> Mg^{2+} | + <u>1</u> SO_3^{2-} | → | <u>MgSO_3</u> | <u>magnesium sulfite</u> |
| (F) | <u>1</u> Cu^{2+} | + <u>2</u> ClO^- | → | <u>$\text{Cu}(\text{ClO})_2$</u> | <u>Copper (II) hypochlorite</u> |
| (G) | <u>1</u> NH_4^+ | + <u>1</u> $\text{C}_2\text{H}_3\text{O}_2^-$ | → | <u>$\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$</u> | <u>ammonium acetate</u> |
| (H) | <u>2</u> Cr^{3+} | + <u>3</u> SO_4^{2-} | → | <u>$\text{Cr}_2(\text{SO}_4)_3$</u> | Chromium (III) sulfate |
| (I) | <u>2</u> Al^{3+} | + <u>3</u> CO_3^{2-} | → | <u>$\text{Al}_2(\text{CO}_3)_3$</u> | <u>aluminum carbonate</u> |
| (J) | <u>1</u> Pb^{2+} | + <u>2</u> OH^- | → | <u>$\text{Pb}(\text{OH})_2$</u> | <u>lead (II) hydroxide</u> |

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M Walker



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