

Moles and calculations

Define moles: It is the amount of substance which is equal to the Avogadro's number of particles.

Avogadro's number or Avogadro's Constant = 6.02×10^{23} particles

Formulas for calculation of moles:

$$1. \quad \text{mole} = \frac{\text{mass}}{\text{Mass of one mole}}$$

Mass of one mole = molar mass = Mr or Ar

$$2. \quad \text{mole of any gas} = \frac{\text{Volume of gas in cm}^3}{\text{Volume of one mole of gas}} = \frac{\text{Volume of gas in dm}^3}{\text{Volume of one mole of gas}}$$

Volume of one mole of any gas = molar gas volume = 24000 cm^3 or 24 dm^3

$$3. \quad \text{mole of any substance} = \frac{\text{volume in dm}^3 \times \text{concentration in mol dm}^{-3}}{\text{(in the form of solution)}}$$

$$\text{Volume in dm}^3 = \frac{\text{Volume of solution in cm}^3}{1000}$$

$$4. \quad \text{percentage yield} = \frac{\text{Experimental yield}}{\text{Theoretical yield}} \times 100$$

$$5. \quad \text{percentage purity} = \frac{\text{Mass of pure substance}}{\text{Mass of impure substance}} \times 100$$

$$6. \quad \text{percentage composition by mass} = \frac{\text{No. of atoms} \times \text{Ar of atom}}{\text{Mass of total compound}} \times 100$$

$$7. \quad \text{molecular formula} = n \times \text{empirical formula}$$

$$\text{Where } n = \frac{\text{Molecular formula mass}}{\text{Empirical formula mass}}$$

Skill Practice 40 Moles And Formula Practice

Ralph Thomas Myers



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