

Name:		
	Date:	

Hour: \_\_\_

#### I. Single Replacement Reactions

(Remember, some reactions will NOT proceed. If there is no reaction, write "NR" to the right of the arrow.)

- Write the reaction for each of the following reactions:
  - a) Sodium sulfate reacts with magnesium to produce magnesium sulfate and sodium
  - b) Copper(I) carbonate reacts with aluminum to form copper metal and aluminum carbonate.
- 2. Complete the following reactions:
  - a) Ca(NO<sub>3</sub>)<sub>2</sub> + AI →
  - b) Zn + K<sub>3</sub>PO<sub>4</sub> →

#### II. Double Replacement Reactions

- 3. Write the reaction for each of the following reactions:
  - a) Iron(II) phosphate reacts with sodium chloride to produce iron(II) chloride and sodium phosphate.
  - b) Calcium carbonate reacts with lithium hydroxide to produce calcium hydroxide and lithium carbonate.
- 4. Complete the following reactions:
  - a) Mg(NO<sub>3</sub>)<sub>2</sub> + Al(C<sub>2</sub>H<sub>3</sub>O<sub>2</sub>)<sub>5</sub> →
  - b) Ba(OH)<sub>2</sub> + Na<sub>3</sub>PO<sub>4</sub> →

#### III. Combustion Reactions

You can try the following formula to balance:  $(CxHy + (x + y/4)O_2 \rightarrow xCO_2 + yH_2O)$ 

- Write the combustion reactions for the following compounds.
  - a) C<sub>2</sub>H<sub>4</sub>
  - b) C<sub>8</sub>H<sub>16</sub>

# **Skill Practice 28 Reactions Answers**

Alfred W. Briggs

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