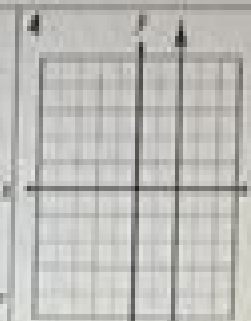
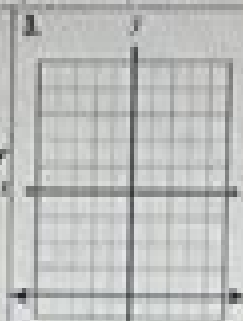
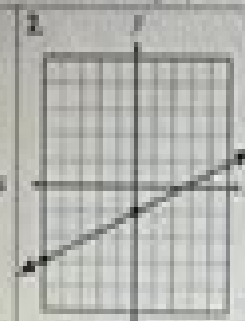
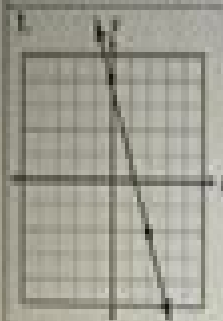


**Directions:** Write the equation of the graph in slope-intercept form.



**Directions:** Convert the equations from standard form to slope-intercept form.

10

1997

100

1. *Journal of the American Medical Association*, 1997; 277: 1001-1005.

1. *Journal of Management Studies*, 1996, 33, 1, 1-15.

10. *Journal of the American Medical Association*, 2000; 283: 2686-2692.

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web during this reaction the  $\text{h}^+$  ions from the acid and the  $\text{oh}^-$  ions from the alkali react together to form water as this happens energy is released which is detected by the thermometer

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exothermic producing up to 235 kcal mole for lithium fluoride the alkali metals react with nonmetals in groups 15 and 16 va  
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measure an appropriate volume of each liquid eg 25 cm<sup>3</sup> place one of the

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web for reactions involving acetic acid or ammonia the measured enthalpy change of neutralization is a few kJ less  
exothermic than with strong acids and bases for example one source which gives the enthalpy change of neutralization of  
sodium hydroxide solution with hcl as 57.9 kJ mol<sup>-1</sup>

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eg 1 hcl aq + naoh aq → nacl aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup> eg 2 h<sub>2</sub>no<sub>3</sub> aq + koh aq → kno<sub>3</sub> aq + h<sub>2</sub>o l for which Δh = 57 kJ mol<sup>-1</sup>

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