

Mole Review Problems

Block _____

Solve the following problems and SHOW YOUR WORK

1. How many moles of ATOMS are there in
- 1.20×10^{25}
- methane (
- CH_4
-) molecules?

$$1.20 \times 10^{25} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) \left(\frac{5 \text{ atoms}}{1 \text{ molecule}} \right) = \underline{99.7 \text{ moles of atoms}}$$

2. Calculate the molarity of a solution formed by dissolving 32.5 g of NaCl in sufficient water to make 2.5 L of solution.

$$\text{FW NaCl} = 23.0 + 35.5 = 58.5 \text{ g/mol}$$

$$32.5 \text{ g} \div 58.5 \text{ g/mol} = 0.556 \text{ mol}$$

$$M = \frac{n}{V} = \frac{0.556 \text{ mol}}{2.5 \text{ L}} = \underline{0.22 \text{ M}}$$

3. How many ATOMS are contained in 32.0 grams of gaseous oxygen?

$$\text{FW O}_2 = (16.0) \times 2 = 32.0 \text{ g/mol}$$

$$32.0 \text{ g} \div 32.0 \text{ g/mol} = 1.00 \text{ mol O}_2$$

$$1.00 \text{ mol O}_2 \times \frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \times \frac{2 \text{ O atoms}}{1 \text{ O}_2 \text{ molecule}} = \underline{1.20 \times 10^{24} \text{ atoms}}$$

4. 0.500 mole of an element has a mass of 31.75 grams. Identify the element.

$$\text{FW} = \frac{31.75 \text{ g}}{0.500 \text{ mol}} = \underline{63.5 \text{ g/mol}} \quad \underline{\text{The element is Copper.}}$$

5. Calculate the number of moles of neon gas in 54 litres at STP.

$$54 \text{ L} \div 22.4 \text{ L/mol} = \underline{2.4 \text{ mol}}$$

6. A sample of carbon reacts with a 9.2 gram sample of hydrogen to form a gaseous compound that has a mass of 120.0 grams. What is the molecular formula of this hydrocarbon, which has a molar mass of 26.0 grams?

$$\text{H: } 9.2 \text{ g} \div 1.0 \text{ g/mol} = 9.2 \text{ moles H}$$

$$\text{C: } (120.0 \text{ g} - 9.2 \text{ g}) \div 12.0 \text{ g/mol} = 9.2 \text{ moles C}$$

$$\text{Mole Ratio } \text{C} : \text{H} = 1 : 1$$

$$\text{Empirical formula is } \text{CH} \rightarrow \text{molar mass } 12.0 + 1.0 = 13.0$$

$$\text{molecular mass is } 26.0 \rightarrow \underline{\text{so formula is } \text{C}_2\text{H}_2}$$

7. What is the mass of
- 2.11×10^{24}
- molecules of chlorine gas?

$$2.11 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = 3.50 \text{ moles Cl}_2$$

$$\text{FW Cl}_2 = 2(35.5) = 71.0 \text{ g/mol}$$

$$3.50 \text{ moles} \times 71.0 \text{ g/mol} = \underline{249 \text{ g}}$$

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Lingsheng Yao



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