

Enthalpy change for the thermal decomposition of potassium hydrogen carbonate =  
 $5.38 \text{ KJ/mol} + 0.32 \text{ KJ/mol} = 5.7 \text{ KJ/mol}$

**Enthalpy change for the thermal decomposition of potassium hydrogen carbonate = 5.7 KJ/mol**

4. When  $\text{KHCO}_3$  is heated, it decomposes according to the following chemical reaction:



Apply Hess' Law and use the results from questions **2b** and **3b** to calculate the enthalpy change for the thermal decomposition of potassium hydrogen carbonate.

# Standard Enthalpy Change Of Potassium Hydrogen Carbonate

**Paddy Gannon**



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