Laboratory experiment 2

Standardization of Silver nitrate solution by Sodium chloride

Objective: to determine the normality (molar concentration of the equivalent) of silver nitrate solution.

Standardization of AgNO3 solution is based on the reaction:

$$AgNO_s + NaCl \leftrightarrow AgCl \downarrow + NaNO_s$$

 $m_e(NaCl) - M(NaCl) \cdot f; \quad f = 1$

<u>Procedure.</u> Completely fill the burette with the solution of silver nitrate AgNO₃ and remove the air from the tip by running out some of the liquid into an empty beaker. Make sure that the lower part of the meniscus is at the zero mark.

Using a pipette, transfer a 20.00 mL aliquot of 0.1000 N solution of sodium chloride (NaCl) to an Erlenmeyer flask and add 2-4 drops of potassium chromate K₂CrO₄ solution (w=5%). Titrate this solution with silver nitrate solution: slowly add the AgNO₃ solution to the flask with sodium chloride solution while gently stirring the contents of the flask. Continue titrating until the color of precipitate changes from white to reddish-brown.

Record the final burette reading. Repeat the titration two more times to a reproducible result. Calculate the normality of the AgNO₃ and correction factor.

$$N(AgNO_1) = \frac{N(NoCI) \cdot V(NoCI)}{V(AgNO_1)} =$$

$$K = \frac{N_{post}}{N_{test}} =$$

Table

Nie	Volume of sodium chloride, mL	Volume of silver nitrate, mL	The mean of volume (AgNO ₃), mL	Normality of silver natrate, mol/L
1	20.00	21.00		
2	20.00	20.90		
3	20.00	20.90		

Conclusion:		

Standardization Of Agno3 Lab Report

Peipei Pang

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- 1. Understanding the eBook Standardization Of Agno3 Lab Report
 - The Rise of Digital Reading Standardization Of Agno3 Lab Report
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Standardization Of Agno3 Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Standardization Of Agno3 Lab Report
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Standardization Of Agno3 Lab Report
 - Personalized Recommendations
 - Standardization Of Agno3 Lab Report User Reviews and Ratings
 - Standardization Of Agno3 Lab Report and Bestseller Lists

- 5. Accessing Standardization Of Agno3 Lab Report Free and Paid eBooks
 - Standardization Of Agno3 Lab Report Public Domain eBooks
 - Standardization Of Agno3 Lab Report eBook Subscription Services
 - Standardization Of Agno3 Lab Report Budget-Friendly Options
- 6. Navigating Standardization Of Agno3 Lab Report eBook Formats
 - ∘ ePub, PDF, MOBI, and More
 - Standardization Of Agno3 Lab Report Compatibility with Devices
 - Standardization Of Agno3 Lab Report Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Standardization Of Agno3 Lab Report
 - Highlighting and Note-Taking Standardization Of Agno3 Lab Report
 - Interactive Elements Standardization Of Agno3 Lab Report
- 8. Staying Engaged with Standardization Of Agno3 Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - $\circ\,$ Following Authors and Publishers Standardization Of Agno3 Lab Report
- 9. Balancing eBooks and Physical Books Standardization Of Agno3 Lab Report
 - Benefits of a Digital Library
 - \circ Creating a Diverse Reading Collection Standardization Of Agno3 Lab Report
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Standardization Of Agno3 Lab Report
 - Setting Reading Goals Standardization Of Agno3 Lab Report
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Standardization Of Agno3 Lab Report
 - Fact-Checking eBook Content of Standardization Of Agno3 Lab Report
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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