

Standardized Test prep Answers

Chapter 1 Page 25

1. C 4. B 7. C
2. C 5. C 8. A
3. D 6. C 9. B

10. Answers may vary.

11. Observations C and G are chemical properties; the others are physical properties.

12. In a chemical change, one or more substances are converted into different substances. A physical change does not involve a change in the identity of the substance or substances present.

13. metals: shiny; good conductors of heat; good conductors of electricity; malleable or ductile; most are solids at room temperature
nonmetals: poor conductors of heat; poor conductors of electricity; many are gases at room temperature; those that are solids are brittle rather than malleable or ductile metalloids: properties intermediate between those of metals and nonmetals; less malleable than metals but not as brittle as solid nonmetals; most are semiconductors of electricity

Chapter 3 Page 93

1. C 4. A 7. B
2. D 5. B 8. D
3. C 6. A 9. C

10. Argon-40 has 22 neutrons ($40 - 18 = 22$), and potassium-40 has 21 neutrons ($40 - 19 = 21$).

11. 6.17 g

12. All cathode rays are the same, regardless of their source. Therefore, the particles responsible for the cathode rays must be present in all atoms. The particles are electrons.

13. When the average atomic mass is calculated, it is 10.811. Because the atomic mass is the same as the atomic mass of boron, mythium was not a new element.

Chapter 5 Page 1

1. C
2. A
3. A
4. A
5. C
6. A
7. D
8. B
9. C

10. S: 2251 kJ/mol; Cl: 2297 kJ/mol; Ar: 2666 kJ/mol; K: 3051 kJ/mol; Ca: 1345 kJ/mol; Sc: 1235 kJ/mol; Ti: 1310 kJ/mol. For the second ionization, the general trend is for increasing IE2 across the period in Groups 2–18 with Group 1 having the highest IE2. IE2 decreases going down a group.

11. Group 16 most commonly forms 2- ions, because these elements require only two more electrons to fill their shell (obtain a noble-gas configuration).

Chapter 2 Page 63

1. B 4. B 7. A
2. C 5. A 8. B
3. C 6. B 9. B

10. 2 cups; direct proportion

11. Statements A and B contain exact numbers.

12. The type of fertilizer is the variable being tested. Control factors are the types of radishes, the amount of water and the amount of sunshine. One control row should be planted under the same control factors but with no fertilizer. There are at least four things that could be used to determine the results: size, quantity, appearance, and taste. Analysis might include bar graphs of each of these measurements for each of the five fertilizer types and the no-fertilizer control row.

13. A unit must be defined in a way that does not depend on the circumstances of the measurement. Not every thumbnail is the same size.

Chap 4 Page 129

1. A 4. D 7. C
2. D 5. B 8. B
3. A 6. B 9. C

10. The color will be orange. Converting energy into frequency gives 4.8×10^{14} , which corresponds to the frequency of orange light.

11. $\frac{1}{3s}$ $\frac{1}{3s}$ $\frac{1}{3p}$ $\frac{1}{3p}$

12. Electrons in atoms can occupy orbitals of only specific energies. When an atom is excited, the electron is no longer in the ground state. When the electron returns to a lower energy level, light is emitted. Because only specific energies are allowed, certain wavelengths of light are emitted, giving rise to the individual lines in the spectrum.

13. Photons of blue light are higher energy than photons of yellow light. Electrons can be emitted only when a photon of sufficient energy strikes the surface of the metal. Therefore, the energy of blue light is greater than the threshold energy, but the energy of yellow light is not.

Chapter 6 Page 215

1. A 2. C 7. A
3. B 4. D 8. B
5. A 6. D 9. C

10. Hybridization explains how the orbitals of an atom become rearranged when the atom forms covalent bonds.

11. Ionic crystals are brittle because shifting of the layers of ions results in large repulsive forces that cause the layers to part completely.

12. Naphthalene has the stronger intermolecular forces even though it is non-polar, because its boiling point is higher than that of acetic acid. Boiling point is directly correlated to strength of intermolecular forces; the stronger the intermolecular forces, the more energy needed to break all the intermolecular forces, and therefore the higher the boiling point. Naphthalene is so large that its dispersion forces are greater than the sum of the dispersion forces and hydrogen bonding in acetic acid.

13. There is a repulsive region at small bond distances, a minimum well at an intermediate bond distance, and an asymptotic region at large bond distances. The repulsive region is due to the repulsion between the two positive nuclei at distances where the nuclei are close to each other. The asymptotic region is where the atoms are far apart and therefore do not interact with each other. The potential energy decreases as the atoms come closer together and are attracted to each other. The minimum well is the region where the attractive forces are balanced by the repulsive forces. This distance is the equilibrium bond length.

Standardized Test Prep Chemistry Answers

Raffaela Di Napoli



Standardized Test Prep Chemistry Answers:

Cracking the SAT II Chemistry Theodore Silver, Princeton Review (Firm), 2005-03 Why The Princeton Review 1 We Know the SAT Chemistry Subject Test The experts at The Princeton Review have spent many years researching the SAT Chemistry Subject Test as well as numerous other standardized tests We re confident this guide delivers the most current and complete information you need to ace this test 2 We Get Results Our inventive approach to standardized test taking has revolutionized the test prep industry and made our courses and tutoring for the SAT and SAT Subject Tests the most popular anywhere The same proven techniques we teach in our courses are also covered in this book 3 We Understand Students Each year we help more than two million students score higher on standardized tests and gain admission to top schools with our books courses tutors and online tools 4 And If It s on the SAT Chemistry Subject Test It s in This Book The Princeton Review realizes that acing the SAT Chemistry Subject Test is very different from getting straight A s in school We don t try to teach you everything there is to know about chemistry only the techniques and information you ll need to maximize your score In *Cracking the SAT Chemistry Subject Test* we ll teach you how to think like the test writers and Master test taking strategies that will improve your score Ace the exam by familiarizing yourself with its format Use Process of Elimination and other proven test taking techniques to solve complicated problems Perfect your test taking skills with practice questions and detailed answers and explanations This book includes three full length practice SAT Chemistry Subject Tests All of our practice test questions are just like those you ll see on the actual test and we fully explain every question Attend Free Practice Tests and Strategy Sessions We re not just good on paper you should see us live The Princeton Review frequently offers free events to students and parents Evaluate Your Options Thousands of students prepare for standardized tests with our books courses and tutoring programs Get on the Inside Track for College Admissions Gaining admission to top colleges takes more than a high test score Other important qualifiers may include a strong admissions essay GPA and volunteer work To learn more about our many books programs and services go to PrincetonReview.com or call us at 800 2Review **The Best Test Preparation for the College Board Achievement Test in Chemistry** Research and Education

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master the test 6 full length practice SAT II Chemistry Subject tests Each test question is answered in complete detail with easy to follow easy to grasp explanations The book s handy Periodic Table of Elements allows for quick answers on the elements appearing on the exam TABLE OF CONTENTS About Research and Education Association Independent Study Schedule CHAPTER 1 ABOUT THE SAT II CHEMISTRY SUBJECT TEST About This Book About The Test How To Use This Book Format of the SAT II Chemistry Scoring the SAT II Chemistry Score Conversion Table Studying for the SAT II Chemistry Test Taking Tips CHAPTER 2 COURSE REVIEW Gases Gas Laws Gas Mixtures and Other Physical Properties of Gases Dalton s Law of Partial Pressures Avogadro s Law The Mole Concept Avogadro s Hypothesis Chemical Compounds and Formulas Mole Concept Molecular Weight and Formula Weight Equivalent Weight Chemical Composition Stoichiometry Weight and Volume Calculations Balancing Chemical Equations Calculations Based on Chemical Equations Limiting Reactant Calculations Solids Phase Diagram Phase Equilibrium Properties of Liquids Density Colligative Properties of Solutions Raoult s Law and Vapor Pressure Osmotic Pressure Solution Chemistry Concentration Units Equilibrium The Law of Mass Action Kinetics and Equilibrium Le Chatelier s Principle and Chemical Equilibrium Acid Base Equilibria Definitions of Acids and Bases Ionization of Water pH Dissociation of Weak Electrolytes Dissociation of Polyprotic Acids Buffers Hydrolysis Thermodynamics I Bond Energies Some Commonly Used Terms in Thermodynamics The First Law of Thermodynamics Enthalpy Hess s Law of Heat Summation Standard States Heat of Vaporization and Heat of Fusion Thermodynamics II Entropy The Second Law of Thermodynamics Standard Entropies and Free Energies Electrochemistry Oxidation and Reduction Electrolytic Cells Non Standard State Cell Potentials Atomic Theory Atomic Weight Types of Bonds Periodic Trends Electronegativity Quantum Chemistry Basic Electron Charges Components of Atomic Structure The Wave Mechanical Model Subshells and Electron Configuration Double and Triple Bonds Organic Chemistry Nomenclature and Structure Alkanes Alkenes Dienes Alkynes Alkyl Halides Cyclic Hydrocarbons Aromatic Hydrocarbons Aryl Halides Ethers and Epoxides Alcohols and Glycols Carboxylic Acids Carboxylic Acid Derivatives Esters Amides Arenes Aldehydes and Ketones Amines Phenols and Quinones Structural Isomerism SIX PRACTICE EXAMS Practice Test 1 Answer Key Detailed Explanations of Answers Practice Test 2 Answer Key Detailed Explanations of Answers Practice Test 3 Answer Key Detailed Explanations of Answers Practice Test 4 Answer Key Detailed Explanations of Answers Practice Test 5 Answer Key Detailed Explanations of Answers Practice Test 6 Answer Key Detailed Explanations of Answers THE PERIODIC TABLE EXCERPT About Research OR 2 To demonstrate proficiency in Chemistry The SAT II Chemistry exam is designed for students who have taken one year of college preparatory chemistry Who Administers The Test The SAT II Chemistry Subject Test is developed by the College Board and administered by Educational Testing Service ETS The test development process involves the assistance of educators throughout the country and is designed and implemented to ensure that the content and difficulty level of the test are appropriate When Should the SAT II Chemistry be Taken If you are applying to a college that requires Subject Test

scores as part of the admissions process you should take the SAT II Chemistry Subject Test toward the end of your junior year or at the beginning of your senior year. If your scores are being used only for placement purposes you may be able to take the test in the spring of your senior year. For more information be sure to contact the colleges to which you are applying.

When and Where is the Test Given? The SAT II Chemistry Subject Test is administered five times a year at many locations throughout the country, mostly high schools. To receive information on upcoming administrations of the exam, consult the publication *Taking the SAT II Subject Tests*, which may be obtained from your guidance counselor or by contacting College Board SAT Program, P.O. Box 6200, Princeton, NJ 08541-6200. Phone: 609 771 7600. Website: <http://www.collegeboard.com>.

Is There a Registration Fee? Yes. There is a registration fee to take the SAT II Chemistry. Consult the publication *Taking the SAT II Subject Tests* for information on the fee structure. Financial assistance may be granted in certain situations. To find out if you qualify and to register for assistance, contact your academic advisor.

HOW TO USE THIS BOOK

What Do I Study First? Remember that the SAT II Chemistry Subject Test is designed to test knowledge that has been acquired throughout your education. Therefore, the best way to prepare for the exam is to refresh yourself by thoroughly studying our review material and taking the sample tests provided in this book. They will familiarize you with the types of questions, directions, and format of the SAT II Chemistry Subject Test. To begin your studies, read over the review and the suggestions for test taking. Take one of the practice tests to determine your areas of weakness and then restudy the review material focusing on your specific problem areas. The course review includes the information you need to know when taking the exam. Be sure to take the remaining practice tests to further test yourself and become familiar with the format of the SAT II Chemistry Subject Test.

When Should I Start Studying? It is never too early to start studying for the SAT II Chemistry test. The earlier you begin, the more time you will have to sharpen your skills. Do not procrastinate. Cramming is not an effective way to study since it does not allow you the time needed to learn the test material. The sooner you learn the format of the exam, the more comfortable you will be when you take the exam.

FORMAT OF THE SAT II CHEMISTRY The SAT II Chemistry is a one-hour exam consisting of 85 multiple-choice questions. The first part of the exam consists of classification questions. This question type presents a list of statements or questions that you must match up with a group of choices lettered A through E. Each choice may be used once, more than once, or not at all. The exam then shifts to relationship analysis questions, which you will answer in a specially numbered section of your answer sheet. You will have to determine if each of two statements is true or false and if the second statement is a correct explanation of the first. The last section is composed strictly of multiple-choice questions with choices lettered A through E.

Material Tested The following chart summarizes the distribution of topics covered on the SAT II Chemistry Subject Test.

Topic	Percentage	Number of Questions
Atomic	10	8
Chemical Equilibrium	10	8
Chemical Kinetics	10	8
Electrochemistry	10	8
Environmental Chemistry	10	8
Organic Chemistry	10	8
Physical Chemistry	10	8
Thermodynamics	10	8
Other	10	8

Atomic, however, do not review too much at any one time. Concentrate on one problem area at a time by reviewing the questions and explanations and by studying our review until you are confident you completely understand the material. Keep track of your scores. By doing so, you will be able to

gauge your progress and discover general weaknesses in particular sections You should carefully study the reviews that cover your areas of difficulty as this will build your skills in those areas

TEST TAKING TIPS

Although you may be unfamiliar with standardized tests such as the SAT II Chemistry Subject Test there are many ways to acquaint yourself with this type of examination and help alleviate your test taking anxieties Become comfortable with the format of the exam When you are practicing to take the SAT II Chemistry Subject Test simulate the conditions under which you will be taking the actual test Stay calm and pace yourself After simulating the test only a couple of times you will boost your chances of doing well and you will be able to sit down for the actual exam with much more confidence Know the directions and format for each section of the test Familiarizing yourself with the directions and format of the exam will not only save you time but will also ensure that you are familiar enough with the SAT II Chemistry Subject Test to avoid nervousness and the mistakes caused by being nervous Do your scratchwork in the margins of the test booklet You will not be given scrap paper during the exam and you may not perform scratchwork on your answer sheet Space is provided in your test booklet to do any necessary work or draw diagrams If you are unsure of an answer guess However if you do guess guess wisely Use the process of elimination by going through each answer to a question and ruling out as many of the answer choices as possible By eliminating three answer choices you give yourself a fifty fifty chance of answering correctly since there will only be two choices left from which to make your guess Mark your answers in the appropriate spaces on the answer sheet Fill in the oval that corresponds to your answer darkly completely and neatly You can change your answer but remember to completely erase your old answer Any stray lines or unnecessary marks may cause the machine to score your answer incorrectly When you have finished working on a section you may want to go back and check to make sure your answers correspond to the correct questions Marking one answer in the wrong space will throw off the rest of your test whether it is graded by machine or by hand You don t have to answer every question You are not penalized if you do not answer every question The only penalty results from answering a question incorrectly Try to use the guessing strategy but if you are truly stumped by a question remember that you do not have to answer it Work quickly and steadily You have a limited amount of time to work on each section so you need to work quickly and steadily Avoid focusing on one problem for too long

Before the Test

Make sure you know where your test center is well in advance of your test day so you do not get lost on the day of the test On the night before the test gather together the materials you will need the next day Your admission ticket Two forms of identification e g driver s license student identification card or current alien registration card Two No 2 pencils with erasers Directions to the test center A watch if you wish but not one that makes noise as it may disturb other test takers On the day of the test you should wake up early after a good night s rest and have breakfast Dress comfortably so that you are not distracted by being too hot or too cold while taking the test Also plan to arrive at the test center early This will allow you to collect your thoughts and relax before the test and will also spare you the stress of being late If you arrive after the test begins you will not be admitted to the test

center and you will not receive a refund During the Test When you arrive at the test center try to find a seat where you feel most comfortable Follow all the rules and instructions given by the test supervisor If you do not you risk being dismissed from the test and having your scores canceled Once all the test materials are passed out the test instructor will give you directions for filling out your answer sheet Fill this sheet out carefully since this information will appear on your score report After the Test When you have completed the SAT II Chemistry Subject Test you may hand in your test materials and leave Then go home and relax When Will I Receive My Score Report and What Will It Look Like You should receive your score report about five weeks after you take the test This report will include your scores percentile ranks and interpretive information

CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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A Bibliography of Standardized Tests for the High School Walter Scott Monroe, 1920 **MCAT General Chemistry**

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