

1. What explains the very high melting and boiling point of water

- |                         |  |                         |  |
|-------------------------|--|-------------------------|--|
| <input type="radio"/> A | Strong dipole-dipole bonds between water molecules | <input type="radio"/> B | Dispersion forces which are present in all molecules |
| <input type="radio"/> C | Strong hydrogen bonds between water molecules      | <input type="radio"/> D | Asymmetrical shape of the polar bonds.               |

2. Does HCl have hydrogen bonding?

- |                         |    |                         |     |
|-------------------------|----|-------------------------|-----|
| <input type="radio"/> A | no | <input type="radio"/> B | yes |
|-------------------------|----|-------------------------|-----|

3. Does CH<sub>4</sub> have hydrogen bonding?

- |                         |     |                         |    |
|-------------------------|-----|-------------------------|----|
| <input type="radio"/> A | yes | <input type="radio"/> B | no |
|-------------------------|-----|-------------------------|----|

4. All molecules have London forces between them, but dipole-dipole and hydrogen bonding are so much stronger that when they are present we can ignore London forces. Which of these has ONLY London forces?

- |                         |                 |                         |                  |
|-------------------------|-----------------|-------------------------|------------------|
| <input type="radio"/> A | I <sub>2</sub>  | <input type="radio"/> B | OCl <sub>2</sub> |
| <input type="radio"/> C | NH <sub>3</sub> | <input type="radio"/> D | SH <sub>2</sub>  |

5. Which substance would have the weakest intermolecular forces of attraction?

- |                         |                 |                         |                  |
|-------------------------|-----------------|-------------------------|------------------|
| <input type="radio"/> A | CH <sub>4</sub> | <input type="radio"/> B | H <sub>2</sub> O |
| <input type="radio"/> C | NaCl            | <input type="radio"/> D | MgF <sub>2</sub> |

6. Intermolecular forces for: NH<sub>3</sub>



- |                         |               |                         |                  |
|-------------------------|---------------|-------------------------|------------------|
| <input type="radio"/> A | Dipole dipole | <input type="radio"/> B | Dispersion Force |
|-------------------------|---------------|-------------------------|------------------|

# States Of Matter And Intermolecular Forces Test

**Daniela Niemeyer**



## States Of Matter And Intermolecular Forces Test:

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Classification of Hydrocarbons Unit XI V Environmental Chemistry Content 1 Some Basic Concepts of Chemistry 2 Structure of Atom 3 Classification of Elements and Periodicity in Properties 4 Chemical Bonding and Molecular Structure 5 States of Matter 6 Thermodynamics 7 Equilibrium 8 Redox Reactions 9 Hydrogen 10 s Block Elements 11 p Block Elements 12 Organic Chemistry Some Basic Principles and Techniques 13 Hydrocarbons 14 Environmental Chemistry I Appendix II Log antilog Table

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