

STOICHIOMETRY NOTES WITH LIMITING REACTANT AND PERCENT YIELD

STOICHIOMETRY

Stoichiometry is the relationship between the masses of the reactants and products of chemical reactions.

- Stoichiometry is possible because of the Law of Conservation of Mass.
- The law of conservation of mass states that the amount of matter before a reaction occurs is equal to the amount of matter present after the reaction occurs.
- Thanks to this idea, we can build ratios between reactants and product that allow us to calculate amounts of each.

Using Ratios

Reactant and product amounts are related in a predictable way.

Each coefficient in the balanced equation can be used to define a ratio between the individual reactants and products of the chemical reaction. For the reaction below, derive only one ratio of choice.

Consider the reaction:

$$2\text{NaCl} + \text{MgSO}_4 \rightarrow \text{Na}_2\text{SO}_4 + \text{MgCl}_2$$

We can write down ratios:

moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	and	$\frac{2\text{NaCl}}{1\text{MgSO}_4}$
moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	and	$\frac{2\text{NaCl}}{2\text{Na}_2\text{SO}_4}$
moles of NaCl	$\frac{2\text{NaCl}}{2\text{NaCl}}$	and	$\frac{2\text{NaCl}}{2\text{MgCl}_2}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$	and	$\frac{1\text{MgSO}_4}{2\text{NaCl}}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$	and	$\frac{1\text{MgSO}_4}{1\text{Na}_2\text{SO}_4}$
moles of MgSO ₄	$\frac{1\text{MgSO}_4}{1\text{MgSO}_4}$	and	$\frac{1\text{MgSO}_4}{2\text{MgCl}_2}$

Example: Finding the Limiting Reactant in the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

5.37 grams of NaCl react with 5.63 grams of Mg. What is the limiting reactant and what mass of MgSO_4 can be formed?

How many moles of NaCl react? $\frac{5.37\text{g NaCl}}{58.44\text{g NaCl/mol}} = 0.0919\text{mol NaCl}$

How many moles of Mg react? $\frac{5.63\text{g Mg}}{24.31\text{g Mg/mol}} = 0.232\text{mol Mg}$

How many moles of MgSO_4 can be formed from the NaCl? $0.0919\text{mol NaCl} \times \frac{1\text{mol MgSO}_4}{2\text{mol NaCl}} = 0.0459\text{mol MgSO}_4$

How many moles of MgSO_4 can be formed from the Mg? $0.232\text{mol Mg} \times \frac{1\text{mol MgSO}_4}{1\text{mol Mg}} = 0.232\text{mol MgSO}_4$

0.0459 moles of MgSO_4 will form and NaCl is the limiting reactant.

Using Mole Ratios

From the last page, we looked at the reaction:

$$2\text{NaCl(aq)} + \text{MgSO}_4\text{(aq)} \rightarrow \text{Na}_2\text{SO}_4\text{(aq)} + \text{MgCl}_2\text{(aq)}$$

We also defined ratios of moles of each substance using the coefficients of the balanced equation.

How do we use the ratios we defined?

How many moles of Mg can be produced from 4 moles of NaCl in the reaction above?

We choose the ratio of moles of Mg to moles of NaCl, remembering to be the given number of moles of NaCl. By writing the ratio first in the denominator, we eliminate the units moles of NaCl and are left with moles of Mg.

$$4\text{ moles NaCl} \times \frac{1\text{ mole Mg}}{2\text{ moles NaCl}} = 2\text{ moles Mg}$$

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web teller s instructions and sample ballot for the meeting of the congregation to call a pastor 1 distribute ballots to all those on active membership roll 2 collect completed ballots 3 go to a secure place to count the ballots 4 the only valid votes are those that have only one mark next to one of the 3 options more than one

church ballots sample slideblast com - May 18 2022

web yes i vote for church name to become a member of church name in mission for christ no i vote for church name not to become a member of church name in mission for christ note this motion must be approved by a simple majority of the votes taken 50 important abstentions are not counted in the calculation

[vahan İsaoglu pastor kut kilisesi linkedin](#) - Mar 16 2022

web kut kilisesi senior pastor İstanbul İstanbul türkiye kut kilisesi kif church mar 2011 halen 12 yıl 6 ay eğitim anadolu Üniversitesi lisans derecesi sosyoloji anadolu Üniversitesi lisans derecesi ekonomi tcm international institute new york city metropol bölgesi burak batmaca