

Chem 1020 Molarity worksheet, version 2

1. What is the molarity of a solution that contains 5.83 g LiCl in 1.50 L solution?

$$5.83 \text{ g LiCl} \times \frac{1 \text{ mol LiCl}}{42.389 \text{ g LiCl}} = 0.13754 \text{ mol LiCl}$$

$$M = \frac{\text{mol solute}}{\text{L solution}} = \frac{0.13754 \text{ mol LiCl}}{1.50 \text{ L solution}} = 0.0917 \text{ M}$$

2. What is the volume of a 0.300 M CuSO₄ solution that contains 12.4 g CuSO₄?

$$12.4 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ L solution}}{0.300 \text{ mol CuSO}_4} = 0.26 \text{ L solution}$$

3. What mass of solid KBr would be needed to prepare 750.0 mL of 1.75 M KBr solution? How specifically would this solution be prepared?

$$750.0 \text{ mL KBr soln} \times \frac{1 \text{ L soln}}{1000 \text{ mL soln}} \times \frac{1.75 \text{ mol KBr}}{1 \text{ L soln}} \times \frac{119.01 \text{ g KBr}}{1 \text{ mol soln}} = 156 \text{ g KBr}$$

Prepare this solution by measuring 156 g KBr into a 750 mL volumetric flask, and adding enough water to make 750 mL solution.

4. What is the % by mass of 28.3 g of solution which contains 3.56 g NaCl?

$$\% \text{ (mass)} = \frac{\text{mass solute}}{\text{mass solution}} \times 100\% = \frac{3.56 \text{ g NaCl}}{28.3 \text{ g solution}} \times 100\% = 12.6\%$$

5. What volume of methanol is required to prepare 2.0 L of a 15.0% by volume solution?

$$\% \text{ (vol)} = \frac{\text{vol solute}}{\text{vol solution}} \times 100\%$$

$$\text{vol solute} = \left(\frac{\% \text{ (vol)}}{100\%} \right) \times \text{vol solution} = \left(\frac{15.0\%}{100\%} \right) \times 2.0 \text{ L} = 0.30 \text{ L}$$

6. What volume of 12.0 M HCl is required to make 500.0 mL of 0.400 M HCl? How specifically would this solution be prepared?

$$M_1 V_1 = M_2 V_2$$

$$V_1 = \frac{M_2 V_2}{M_1} = \frac{(0.400 \text{ M})(500.0 \text{ mL})}{(12.0 \text{ M})} = 16.7 \text{ mL}$$

Prepare this solution by measuring 16.7 mL of 12.0 M HCl into a 500 mL volumetric flask, and adding water to make 500 mL of solution.

7. If 17.0 mL of 16.0 M H₂SO₄ solution is diluted to 100.0 mL, what is the diluted concentration?

$$M_1 V_1 = M_2 V_2$$

$$M_2 = \frac{M_1 V_1}{V_2} = \frac{(16.0 \text{ M})(17.0 \text{ mL})}{(100.0 \text{ mL})} = 2.72 \text{ M}$$

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General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value *Mathematics for the Clinical Laboratory E-Book* Lorraine J. Doucette, 2020-06-18 Selected for Doody's Core Titles 2024 with Essential Purchase designation in Laboratory Technology Master the skills you'll need to perform accurate clinical laboratory calculations Mathematics for the Clinical Laboratory 4th Edition demonstrates the calculations used in the analysis of test specimens It begins by explaining basic mathematical principles and then covers the types of calculations needed in specific areas of the clinical lab including urinalysis hematology and microbiology Finally it focuses on the statistical calculations used in quality assurance and quality control Step by step examples reinforce your understanding and calculation templates and practice problems ensure that you make correct calculations every time Step by step examples explain basic mathematical principles and show you exactly how to perform each type of calculation Sample problems with answers can also be used as templates for solving laboratory calculations Practice problems at the end of each chapter provide a self assessment tool helping you determine what you need to review Summaries of important formulas are included at the end of the text's major sections Coverage of statistical calculations includes standard deviation as well as calculations associated with quality assurance and quality control Quick tips and notes make it easier to understand and remember pertinent information Learning objectives at the beginning of each chapter provide measurable outcomes to achieve by completing the chapter material Full color design includes 100 illustrations Useful appendix of Greek symbols provides a quick reference to turn to when studying Glossary at the back of the textbook includes definitions of important mathematical terms New Updated content and calculations reflect the latest procedures used in today's laboratories

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Resources in Education ,1980-06 *General Chemistry* James P. Birk,Donald M. Baird,1994-03 *The Software Encyclopedia 2000* Bowker Editorial Staff,2000-05

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