TYPES OF CHEMICAL REACTIONS

Complete each of the following equations as needed to make it the type of reaction indicated. Be sure to write each formula correctly.

- Double replacement: Na₂CO₄ + PbCl 2 NaCl + PbCO₄
- 2) Single replacement: Ch + 2 NaBr -> 2 NaCl + Br₂

answer key

- 3) Synthesis: $2 H_2 + 0_2 2 HO$
- 4) Double replacement: 3 Ca(OH)₂ + 2 FeCl 2 Fe(OH)₂ + 3 CaCl₂
- 5) Double replacement: AgNO₃ + KI -> AgI + KNO₃
- 6) Single replacement: Cu + 2 AgNO₃ -> 2 Ag + Cu(NO₃)₂ [Copper (II) is used here]
- 7) Decomposition: Mg(ClO₃)₂ MgCl₂ + 3 0₂
- 8) Single replacement: Fe + Cu(NO₃)₂ -> Fe(NO₃)₂ + Cu [Assume Fe^{2*}]
- 9) Synthesis: S+02 SO2
- 10) Decomposition: 2 Hg0 4 Hg + 0

For each of the following equations, identify what kind of reaction it represents: double replacement, single replacement, decomposition, or synthesis (composition).

- 1) 2Mg + 0₂ 2 MgO Synthesis
- 2) Cal₂ + Cl₂ CaCl₂ In Single Replacement
- 3) 3KOH + AICl₃ -> AK(OH)₃ + 3KCl Double Replacement
- 4) C + O2 -> CO2 Synthesis

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H Kauffman

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