

Chemistry – Unit 4 Worksheet 2

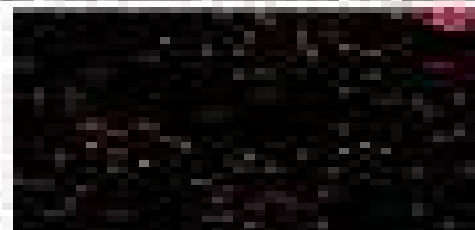
Rutherford's Experiment

An alpha (α) particle (essentially a He^{2+} ion) is positively charged and is much heavier than electrons (e^{-}). It is, therefore, much less deflected by the electric fields of the electrons. Alpha particles are emitted from certain unstable isotopes of elements and they pass through thin sheets of metal foil (usually a few millimetres thick) that support a screen at the end of a 'fluorescent tube'.

Part 1

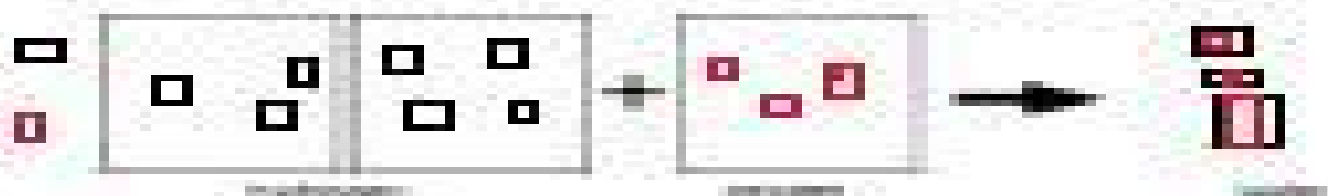
- What is a cathode ray? Describe where the source of your particles is positioned in the experiment and how you observe the results of the experiment.

- Describe the historical idea that cathode rays were the same as the particles.



Rutherford used a fluorescent substance coating the end of the tube, after first sealing the gases out of the system. The α particles ionized the gas, causing the fluorescence of the gas. It was possible to observe that the results of Rutherford's experiment refuted the previous model.

- How does your observation of a gas, releasing of light upon gas being subjected to a vacuum of hydrogen gas fit the model? What is the evidence that the cathode ray is not hydrogen ions? Is it possible that the cathode ray is made up of hydrogen ions? Is it possible that the cathode ray is made up of gas molecules that have ionized? Is possible?
- Describe the structure of hydrogen and oxygen in the atomistic model. How does the atomistic model for the system contradict the results of the experiment?



U4 Ws2 V2 Chemistry Answers

Bev Cooper, Graeme Abbott



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