Chemistry - Unit 6 Reaction Equations Worksheet 1

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

- 1. ___C + ___H₂O \rightarrow ___CO + ___H₂
- 2. $\underline{\hspace{0.1cm}}$ MgO \rightarrow $\underline{\hspace{0.1cm}}$ Mg + $\underline{\hspace{0.1cm}}$ O₂
- 3. $Al + O_2 \rightarrow Al_2O_3$

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Before

#8

- #3 Before
- 5. ___Cl₂ + ___KI → ___KCl + ___I₂

 $_{\text{Zn}}$ + $_{\text{H}_2}$ SO₄ \rightarrow $_{\text{ZnSO}_4}$ + $_{\text{H}_2}$

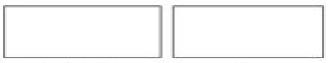
- CuCl → Cu + Cl₂
- ___Na + ___Cl₂ → ___NaCl
- 8. __Al + __HCl \rightarrow __AlCl₃ + __H₂



After

After

- 9. $_{\text{Fe}_2O_3} \rightarrow _{\text{Fe}} + _{\text{O}_2}$
- 10. ___P + ___O₂ \rightarrow ___P₂O₅
- 11. $\underline{\hspace{0.5cm}}Mg + \underline{\hspace{0.5cm}}HCl \rightarrow \underline{\hspace{0.5cm}}MgCl_{2} + \underline{\hspace{0.5cm}}H_{2}$
- 12. $\underline{\hspace{1cm}} H_2 + \underline{\hspace{1cm}} N_2 \rightarrow \underline{\hspace{1cm}} NH_3$
- 13. $__BaCl_2 + __H_2SO_4 \rightarrow __BaSO_4 + __HCl$
- 14. $__{CH_4} + __{O_2} \rightarrow __{CO_2} + __{H_2O}$



- #14 Before After
 15. a) ZnCl. + (NH.)-S → ZnS + NH.Cl
 - b) Find the molar mass of these reactants.
 - c) How many moles of ZnCl₂ would be in 25 g? How much mass would 0.55 moles of (NH₄)₂S have?

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Peipei Pang

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Table of Contents U6 Ws1 V2 Answers

- 1. Understanding the eBook U6 Ws1 V2 Answers
 - The Rise of Digital Reading U6 Ws1 V2 Answers
 - Advantages of eBooks Over Traditional Books
- 2. Identifying U6 Ws1 V2 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an U6 Ws1 V2 Answers
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from U6 Ws1 V2 Answers
 - Personalized Recommendations
 - U6 Ws1 V2 Answers User Reviews and Ratings

- U6 Ws1 V2 Answers and Bestseller Lists
- 5. Accessing U6 Ws1 V2 Answers Free and Paid eBooks
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 - U6 Ws1 V2 Answers eBook Subscription Services
 - U6 Ws1 V2 Answers Budget-Friendly Options
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 - o ePub, PDF, MOBI, and More
 - ∘ U6 Ws1 V2 Answers Compatibility with Devices
 - U6 Ws1 V2 Answers Enhanced eBook Features
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 - Highlighting and Note-Taking U6 Ws1 V2 Answers
 - Interactive Elements U6 Ws1 V2 Answers
- 8. Staying Engaged with U6 Ws1 V2 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers U6 Ws1 V2 Answers
- 9. Balancing eBooks and Physical Books U6 Ws1 V2 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection U6 Ws1 V2 Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine U6 Ws1 V2 Answers
 - Setting Reading Goals U6 Ws1 V2 Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of U6 Ws1 V2 Answers
 - Fact-Checking eBook Content of U6 Ws1 V2 Answers
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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