

## Teoria atomica e leggi quantitative

L'Universo è composto di materia ed energia, due aspetti della medesima entità visto che, come dimostrò Einstein, esse possono convertirsi l'una nell'altra secondo la relazione  $E = mc^2$ .

La chimica è la scienza che studia le caratteristiche, la struttura e le trasformazioni della materia e gli scambi energetici connessi a tali trasformazioni. Oggi noi possediamo un modello sufficientemente dettagliato per descrivere la struttura della materia, il modello atomico. Parlare di atomi è attualmente quasi un fatto scontato, ma il modello atomico si affermò con difficoltà, nonostante sia stato proposto per la prima volta più di duemila anni fa. Le prime ipotesi sulla struttura e sulla natura della materia vennero infatti avanzate, agli albori del pensiero occidentale, in Grecia. A tal proposito possiamo individuare, in estrema sintesi, due problemi fondamentali intorno ai quali si sviluppò il dibattito filosofico su questo tema: A) continuità/discontinuità della materia: l'atomo Da una parte Aristotele che ipotizza l'esistenza di una materia continua, divisibile indefinitamente in frammenti sempre più piccoli e quindi irriducibile ad unità elementari, due correnti di pensiero che interpretavano la materia in modi diametralmente opposti. Dall'altra Democrito di Abdera (IV sec. a.C.), anche se il primo ad avere parlato di atomi fu Leucippo, che ritiene invece la materia discontinua e sostiene pertanto l'esistenza di particelle minuscole, chiamate atomi<sup>1</sup>, invisibili, incorruttibili ed eterne. Le idee di Democrito sopravvissero nei secoli. Furono divulgate da Epicuro di Samo (III sec. a.C.) e, nel mondo latino, da Tito Lucrezio Caro (95-55 a.C.) nel "De Rerum Natura". B) complessità/semplività della materia: l'elemento La ricerca dell'arché, del 'principio primo', dal quale discendeva tutta la molteplicità dell'essere, rappresentò forse l'elemento peculiare della prima speculazione greca intorno al mondo. Il problema di poter ricondurre l'enorme varietà di sostanze conosciute alla combinazione di poche sostanze semplici significava fondamentalmente razionalizzare il mondo e quindi "spiegare" i fenomeni complessi riconducendoli e riducendoli alla loro composizione elementare. A parte i primi tentativi effettuati in tal senso dai primi 'Fisiologi' presocratici, come Talete (624-546 a.C.) che individuò nell'acqua il principio primo o Anassimene (586-528 a.C.) che lo identificò nell'aria, l'ipotesi che ebbe la maggior fortuna durante tutto il medioevo, grazie all'autorevole consenso di Aristotele, fu senza dubbio quella di Empedocle (490-430 a.C.). Secondo tale ipotesi tutta la materia era composta da quattro sostanze fondamentali ed elementari: l'aria, l'acqua, la terra e il fuoco. Il concetto di atomo e di elemento sono fondamentali in chimica. Le prime definizioni moderne si devono a Boyle e a Newton. Per R.Boyle (1627-1691) gli elementi sono "corpi primitivi, semplici, incontaminati, che, non essendo costituiti da altri corpi o di loro mescolanze, sono ingredienti di cui i corpi misti sono costituiti e nei quali questi possono essere in definitiva risolti".

Mentre Newton immaginò gli atomi come minuscole sfere, dominate solo da forze attrattive e repulsive. In Opticks così scrive: "In principio Dio creò la materia in particelle mobili, impenetrabili, dure, massicce, solide,..." Oggi la nostra idea di atomo è notevolmente cambiata rispetto al modello newtoniano, ciononostante vi sono ambiti della chimica e della fisica moderna (teoria cinetica dei gas, ad esempio) in cui tale modello risulta essere ancora perfettamente adeguato ed in grado di giustificare alcuni comportamenti della materia. La chimica moderna, che nasce tra la fine del '700 e l'inizio dell'800, giunge ad unificare il concetto di atomo e di elemento. Tale unificazione si produce con la formulazione da parte di Dalton della Teoria atomica, che possiamo così sintetizzare - La materia è composta di atomi indivisibili ed inalterabili; - Esistono atomi con caratteristiche differenti. Tutti gli atomi di uno stesso tipo costituiscono le sostanze semplici o elementi; - Le trasformazioni chimiche si producono per unione o separazione di atomi secondo rapporti rigorosamente determinati e caratteristici per ogni sostanza chimica. La teoria atomica è stata in seguito sostanzialmente confermata ed è a tutt'oggi uno dei paradigmi della chimica. In natura esistono 90 tipi di atomi che individuano altrettanti elementi chimici (gli elementi fino all'Uranio sono 92, ma il 43mo (Tc Tecnezio) ed il 61mo (Pm Promezio) sono prodotti artificialmente). Tuttavia oggi gli atomi non sono più considerati strutture elementari ed "indivisibili". Essi sono infatti composti da tre tipi di particelle subatomiche: il protone, il neutrone e l'elettrone. Il protone, con carica elettrica positiva, ed il neutrone,

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