

The Main Types of Chemical Reactions

A chemical reaction may be regarded as the process that occurs when matter undergoes change in composition. To be able to understand chemical reactions, one needs to recognize the key types of reactions. Reactions may be classified in several different ways, namely:

Synthesis (Combination) Reactions; Decomposition Reactions; Substitution (single replacement) Reactions; Precipitation (Double replacement) Reactions; Neutralisation (acid-base) Reactions; Redox (Reduction and Oxidation) Reactions.

Type of Reaction	Explanation	General Equation (General form of change occurrences)	Example
Synthesis/ Combination	Reaction in which two or more simpler substances (elements) combine chemically to give a compound.	$A + B \rightarrow AB$	$2H_2 + O_2 \rightarrow 2H_2O$
Decomposition	Reaction in which a compound is broken down into simpler substances (elements).	$AB \rightarrow A + B$	$2H_2O \rightarrow 2H_2 + O_2$
Substitution (Single replacement)	Reaction in which an atom or group of atoms is replaced by another atom or group.	$A + BC \rightarrow AC + B$ OR $A + BC \rightarrow BA + C$	$Zn + 2HCl \rightarrow ZnCl_2 + H_2$ OR $Cl_2 + 2NaBr \rightarrow 2NaCl + Br_2$
Precipitation (Double replacement)	Reaction in which a solid compound is formed when solutions of two soluble compounds are mixed.	$AB + CD \rightarrow AD + CB$	$AgNO_3 + NaCl \rightarrow AgCl + NaNO_3$
Neutralisation (Acid- Base)	Reaction in which an acid reacts with a base to give salt and water.	$HA + BOH \rightarrow H_2O + BA$	$HCl + NaOH \rightarrow H_2O + NaCl$
Redox (Reduction- Oxidation)	Reaction in which electron transfer occurs.	Reduction: $A \rightarrow A + e^-$ Oxidation: $B + e^- \rightarrow B^-$	$aCl_2 + 2e^- \rightarrow 2Cl^-$ (gain electrons) $Na \rightarrow Na^+ + e^-$ (loss of electrons) $2Na + Cl_2 \rightarrow 2 NaCl$

Types Of Chemical Reaction Answers

SA Adler



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Chemical Misconceptions Keith Taber, 2002 Part 2 provides strategies for dealing with some of the misconceptions that students have by including ready to use classroom resources **Educart NCERT Exemplar Class 10 Science 2025**

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This text has been specifically designed to prepare people with previously limited chemical knowledge for entrance into science related courses such as Foundation and Access courses which involve chemistry in higher education Until now there have been no texts available for use on these courses and this book fills that gap Access to Chemistry effectively forms a self study course which is split into separate modules and units covering the full spread of concepts required for those needing a basic knowledge of chemistry The material is presented in a friendly and easy to use manner which allows the student to pace their acquisition of knowledge and gain increasing confidence in order to succeed in understanding essential relevant concepts Other useful features of this book include starter diagnostic tests worked examples and self study tests with answers at the end of each unit In addition to Access or Foundation course students and their tutors to whom this book will prove essential it will have an appeal also as a revision text for those needing a refresher after a break in the subject In addition it will be of interest to members of the general public who wish to better educate themselves on chemical matters as it provides a clear and useful insight into areas such as health home chemicals business market trends and gardening

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Chemical Reactions, Materials and Particles Gill

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