

CHEMISTRY 11 – WORD EQUATIONS AND TYPES OF CHEMICAL REACTIONS WORKSHEET

- 1) Translate the following word equations into chemical symbols and balance the resulting equations. Do not include the phases.

(a) Potassium + water → potassium hydroxide + hydrogen



(b) Nitrogen + hydrogen → ammonia



- 2) In each of the following write a balanced chemical equation, including the phases.

(a) Liquid phosphoric acid reacts with aqueous barium hydroxide to give water and a precipitate of barium phosphate



(b) Solid aluminium oxide and aqueous sulphuric acid produce water and aqueous aluminium sulphate



(c) Nitrogen trifluoride gas and hydrogen gas react to form nitrogen gas and gaseous hydrogen fluoride



(d) Powdered sodium carbonate and aqueous hydrobromic acid react to form carbon dioxide gas, aqueous sodium bromide and water



- 3) Classify the following reactions.

Reaction	Type
$2\text{C}_2\text{H}_2 + 5\text{O}_2 \rightarrow 4\text{CO}_2 + 2\text{H}_2\text{O}$	Combustion
$\text{Mg} + \text{CuSO}_4 \rightarrow \text{MgSO}_4 + \text{Cu}$	Single Replacement
$4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$	Synthesis
$\text{Sn}(\text{OH})_4 + 4\text{HBr} \rightarrow 4\text{H}_2\text{O} + \text{SnBr}_4$	Neutralization
$2\text{Fe}(\text{NO}_3)_3 + 3\text{MgS} \rightarrow \text{Fe}_2\text{S}_3 + 3\text{Mg}(\text{NO}_3)_2$	Double Replacement
$2\text{N}_2\text{O} \rightarrow 2\text{N}_2 + \text{O}_2$	Decomposition

Types Of Chemical Reactions Lab 14 Equations

Ying Liu



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Table of Contents Types Of Chemical Reactions Lab 14 Equations

1. Understanding the eBook Types Of Chemical Reactions Lab 14 Equations
 - The Rise of Digital Reading Types Of Chemical Reactions Lab 14 Equations
 - Advantages of eBooks Over Traditional Books
2. Identifying Types Of Chemical Reactions Lab 14 Equations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Types Of Chemical Reactions Lab 14 Equations
 - User-Friendly Interface
4. Exploring eBook Recommendations from Types Of Chemical Reactions Lab 14 Equations
 - Personalized Recommendations
 - Types Of Chemical Reactions Lab 14 Equations User Reviews and Ratings
 - Types Of Chemical Reactions Lab 14 Equations and Bestseller Lists

5. Accessing Types Of Chemical Reactions Lab 14 Equations Free and Paid eBooks
 - Types Of Chemical Reactions Lab 14 Equations Public Domain eBooks
 - Types Of Chemical Reactions Lab 14 Equations eBook Subscription Services
 - Types Of Chemical Reactions Lab 14 Equations Budget-Friendly Options
6. Navigating Types Of Chemical Reactions Lab 14 Equations eBook Formats
 - ePub, PDF, MOBI, and More
 - Types Of Chemical Reactions Lab 14 Equations Compatibility with Devices
 - Types Of Chemical Reactions Lab 14 Equations Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Types Of Chemical Reactions Lab 14 Equations
 - Highlighting and Note-Taking Types Of Chemical Reactions Lab 14 Equations
 - Interactive Elements Types Of Chemical Reactions Lab 14 Equations
8. Staying Engaged with Types Of Chemical Reactions Lab 14 Equations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Types Of Chemical Reactions Lab 14 Equations
9. Balancing eBooks and Physical Books Types Of Chemical Reactions Lab 14 Equations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Types Of Chemical Reactions Lab 14 Equations
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Types Of Chemical Reactions Lab 14 Equations
 - Setting Reading Goals Types Of Chemical Reactions Lab 14 Equations
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Types Of Chemical Reactions Lab 14 Equations
 - Fact-Checking eBook Content of Types Of Chemical Reactions Lab 14 Equations
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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