

Lab - Types of Chemical Reactions 4352 (35 points)

Introduction

There are many kinds of chemical reactions and several ways to classify them. One useful method classifies reactions into four major types. These are (1) direct combination, or synthesis; (2) decomposition, or analysis; (3) single replacement; and (4) exchange of ions, or double replacement. Not all reactions can be put into one of these categories. Many, however, can.

In a synthesis reaction, two or more substances (elements or compounds) combine to form a more complex substance. Equations for synthesis reactions have the general form $A + B \rightarrow AB$. For example, the formation of water from hydrogen and oxygen is written $2H_2 + O_2 \rightarrow 2H_2O$.

A decomposition reaction is the opposite of a synthesis reaction. In decomposition, a compound breaks down into two or more simpler substances (elements or compounds). Equations for decomposition reactions have the form $AB \rightarrow A + B$. The breakdown of water into its elements is an example of such a reaction: $2H_2O \rightarrow 2H_2 + O_2$.

In a single replacement reaction, one substance in a compound is replaced by another, more active substance (an element). Equations for single replacement reactions have two general forms. In reactions in which one metal replaces another metal, the general equations is $X + YB \rightarrow XB + Y$. In those in which one nonmetal replaces another nonmetal, the general form is $X + AY \rightarrow AX + Y$. The following equations illustrate these types of reactions:

Zinc metal replaces copper(II) ion:



Chlorine (a nonmetal) replaces bromide ions: $Cl_2(g) + 2KBr(aq) \rightarrow 2KCl(aq) + Br_2(l)$

In a double replacement reaction, the metal ions of two different ionic compounds can be thought of as "replacing one another." Equations for this type of reaction have the general form $AB + CD \rightarrow AD + CB$. Most replacement reactions, both single and double, take place in aqueous solutions containing free ions. In a double replacement reaction, one of the products is a precipitate, an insoluble gas, or water. An example is the reaction between silver nitrate and sodium chloride in which the precipitate silver chloride is formed:



All of the types of reactions discussed here may be represented by balanced molecular equations. Reactions involving ion exchanges may be represented by ionic equations also. In this investigation you will be concerned only with molecular formulas and equations. In a balanced equation, the number of atoms of any given element must be the same on both sides of the equation. Multiplying the coefficient and the subscript of an element must yield the same result on both sides of the balanced equation.

In this investigation you will observe examples of the four types of reaction described above. You will be expected to balance the equations representing the observed reactions.

Types Of Chemical Reactions Lab 14

James McKeen Cattell, Jaques Cattell



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