

**Balancing a Redox Equation for an Acidic Solution
Using the Oxidation Number Method**
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



What Is Required?

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

What Is Given?

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

Plan Your Strategy	Act on Your Strategy
Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.	$\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ <p>The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p>
Step 2 Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number.	<p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the O_2 lost 12 electrons and the CS_2 gained 4 electrons.</p> $ \begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array} $

Solutions Manual Teacher Redox

SJ Ball



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AAAS Science Book List, 1978-1986 Kathryn Wolff, Susan M. O'Connell, Valerie J. Montenegro, 1986 A selected and annotated list of science and mathematics books which supplements the AAAS science book list 3rd ed 1970 and the AAAS science book list supplement 1978

Inorganic Chemistry Gary Wulfsberg, 2000-03-16 Both elementary inorganic reaction

chemistry and more advanced inorganic theories are presented in this one textbook while showing the relationships between the two

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