

Examen de révision de chimie des solutions

Exercice 1

Compléter le tableau suivant :

Solution Aqueuse	PKa	Co en mol/l	[H ₃ O ⁺] en mol/l	[HO ⁻] en mol/l	[Acide] en mol/l	[Base] en mol/l	pH	Coefficient d'ionisation	Type de Solution
Acide		0,001					3		
HCN	9,4	Décimolaire							
NaOH		Molaire							
CH ₃ NH ₂	10,6	Centimolaire							

Exercice 2

On considère une solution S1 d'acide faible HBO₂ de concentration C₁=10⁻²M.

1. Calculer le pH de cette solution (Démonstration)
2. Calculer le coefficient de dissociation

A 1 litre de la solution de HBO₂ précédente, on ajoute 1 litre d'une solution d'acide fort HCl de concentration C₂=10⁻²M.

3. Donner l'expression de H₃O⁺ de la solution obtenue.
4. Calculer le pH de cette solution

A litre de la solution S1 de HBO₂, on ajoute sans variation, 10g du sel NaBO₂, totalement soluble dans l'eau.

5. Ecrire les réactions qui ont lieu dans la solution.
6. Quelle est la nature de la solution obtenue.
7. Calculer le Ph.

pKa (HBO₂/BO₂⁻) = 9,16 ; Masse molaire de NaBO₂ est 66 g/mol

Exercice 3

1. Déterminer la solubilité en mol/l des composés suivants :

Sel	CaSO ₄	AgBr	PbSO ₄	AgCl
Produit de solubilité (Ks)	2,0 10 ⁻⁸	7,7 10 ⁻¹³	2,0 10 ⁻⁸	1,6 10 ⁻¹⁰

Y aura-t-il précipitation :

2. Si On mélange 100 ml de CaCl₂ 0,02 M avec 100 ml de Na₂SO₄ 4.10⁻⁴ M.
3. Si on mélange 250 ml d'une solution de Pb(NO₃)₂ 1,6 10⁻³ M avec 750 ml d'une solution de Na₂SO₄ 2,4 10⁻³ M.
4. On ajoute une solution de nitrate d'argent concentrée goutte à goutte à une solution aqueuse dont la concentration en ions Cl⁻ est de 0,1M et la concentration en ions Br⁻ est de 10⁻⁴ M. Quelle est le sel qui précipité le premier AgCl ou AgBr ??

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