

1 25.0 cm<sup>3</sup> of 0.200 mol/dm<sup>3</sup> sulfuric acid neutralises 18.6 cm<sup>3</sup> of barium by droxide solution.

H<sub>2</sub>SO<sub>\*</sub>(aq) + Ba(OHO<sub>2</sub>(aq) → BaSO<sub>\*</sub>(s) + 2H<sub>2</sub>O(t)

a Find the concentration of the barium hydroxide solution in mol/dm2. Give your answer to 3 significant figures.

moles  $H_1SO_4 = \text{conc} \times \text{vol} (\text{dm}^2) = 0.200 \times \frac{15.8}{1848} = 0.00500 \text{ mol}$ moles  $Ba(OH)_2 = \text{moles of } H_2SO_4 = 0.00500 \text{ mol}$  $\text{conc } Ba(OH)_2 = \frac{\text{moles}}{\text{volume} (\text{dm}^2)} = \frac{0.00500}{1848} = 0.269 \text{ mol} \text{dm}^2$ 

b Find the concentration of the barium hydraxide solution in g/dm³. Give your answer to 3 significant figures.

conc Ba(OH) - 171 x 0.269 - 46,0 g/dm<sup>2</sup>

c List the key steps in the practical you would carry out to find out that 18.6 cm² of the barium hydroxide solution is needed to neutralise 25.0 cm² of 0.200 movidm² sulfuric acid.

use a pipette to place 25 cm² of sulfuric acid in a conical flask add an indicator add barium hydroxide from burette until colour changes go dropwise near end-point record result

2 What volume of 0.0100 mol/dm<sup>3</sup> calcium hydroxide neutralises 25.0 cm<sup>3</sup> of 0.140 mol/dm<sup>3</sup> sulfuric acid? Give your answer to 3 significant figures

H<sub>2</sub>SO<sub>4</sub>(aq) + Ca(OH)<sub>2</sub>(aq) → CaSO<sub>4</sub>(s) + 2 H<sub>2</sub>O<sub>3</sub>()
moles H<sub>2</sub>SO<sub>4</sub> = conc × vol (dm<sup>2</sup>) = 0.140 × 15.8 / 1846 = 0.00350 mol
moles Ca(OH)<sub>2</sub> = moles H<sub>2</sub>SO<sub>4</sub> = 0.00350 mol
volume Ca(OH)<sub>2</sub> = moles / 0.00350 = 0.350 dm<sup>2</sup>

3 What volume of 1.50 mol/dm<sup>3</sup> potassium hydroxide reacts with 25.0 cm<sup>3</sup> of 1.10 mol/dm<sup>3</sup> nitric acid? Give your answer to 3 significant figures

# **Titration Chem Study Guide Answers**

**Christopher Talbot, Richard Harwood** 

## **Titration Chem Study Guide Answers:**

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