

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES  
(CHEMISTRY)**

**TABLE 1      PHYSICAL CONSTANTS**

NAME	SYMBOL	VALUE
Magnitude of charge on electron	$e$	$1,6 \times 10^{-19} \text{ C}$
Mass of an electron	$m_e$	$9,1 \times 10^{-31} \text{ kg}$
Standard pressure	$p^\circ$	$1,01 \times 10^5 \text{ Pa}$
Molar gas volume at STP	$V_m$	$22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$
Standard temperature	$T^\circ$	$273 \text{ K}$
Avogadro's constant	$N_A$	$6,02 \times 10^{23} \text{ mol}^{-1}$
Faraday's constant	$F$	$96\,500 \text{ C} \cdot \text{mol}^{-1}$

**TABLE 2      CHEMISTRY FORMULAE**

$n = \frac{m}{M}$	$n = \frac{N}{N_A}$	$n = \frac{V}{V_m}$
$c = \frac{n}{V}$ OR $c = \frac{m}{MV}$	$K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at $25^\circ \text{C}$ (298 K)	
$q = It$ $q = nF$	$E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$ $E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$	

# Tmrf Exemplar Physical Science P2

**Bill Tillery**

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## **Table of Contents Tmrf Exemplar Physical Science P2**

1. Understanding the eBook Tmrf Exemplar Physical Science P2
  - The Rise of Digital Reading Tmrf Exemplar Physical Science P2
  - Advantages of eBooks Over Traditional Books
2. Identifying Tmrf Exemplar Physical Science P2
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Tmrf Exemplar Physical Science P2
  - User-Friendly Interface
4. Exploring eBook Recommendations from Tmrf Exemplar Physical Science P2
  - Personalized Recommendations
  - Tmrf Exemplar Physical Science P2 User Reviews and Ratings
  - Tmrf Exemplar Physical Science P2 and Bestseller Lists
5. Accessing Tmrf Exemplar Physical Science P2 Free and Paid eBooks
  - Tmrf Exemplar Physical Science P2 Public Domain eBooks
  - Tmrf Exemplar Physical Science P2 eBook Subscription Services
  - Tmrf Exemplar Physical Science P2 Budget-Friendly Options

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    - ePub, PDF, MOBI, and More
    - Tmrf Exemplar Physical Science P2 Compatibility with Devices
    - Tmrf Exemplar Physical Science P2 Enhanced eBook Features
  7. Enhancing Your Reading Experience
    - Adjustable Fonts and Text Sizes of Tmrf Exemplar Physical Science P2
    - Highlighting and Note-Taking Tmrf Exemplar Physical Science P2
    - Interactive Elements Tmrf Exemplar Physical Science P2
  8. Staying Engaged with Tmrf Exemplar Physical Science P2
    - Joining Online Reading Communities
    - Participating in Virtual Book Clubs
    - Following Authors and Publishers Tmrf Exemplar Physical Science P2
  9. Balancing eBooks and Physical Books Tmrf Exemplar Physical Science P2
    - Benefits of a Digital Library
    - Creating a Diverse Reading Collection Tmrf Exemplar Physical Science P2
  10. Overcoming Reading Challenges
    - Dealing with Digital Eye Strain
    - Minimizing Distractions
    - Managing Screen Time
  11. Cultivating a Reading Routine Tmrf Exemplar Physical Science P2
    - Setting Reading Goals Tmrf Exemplar Physical Science P2
    - Carving Out Dedicated Reading Time
  12. Sourcing Reliable Information of Tmrf Exemplar Physical Science P2
    - Fact-Checking eBook Content of Tmrf Exemplar Physical Science P2
    - Distinguishing Credible Sources
  13. Promoting Lifelong Learning
    - Utilizing eBooks for Skill Development
    - Exploring Educational eBooks
  14. Embracing eBook Trends
    - Integration of Multimedia Elements

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