Activity B:	Get the Gizmo ready:	5.3
Determining concentration	 Click Reset. Select 1.00 M NaOH for the Burette and Mystery H₂SO₄ for the Flask. Select Bromthymol blue for the Indicator. 	

Introduction: Adding a drop of strong acid or base into a neutralized solution is similar to adding a drop of strong acid or base to water—it causes an abrupt change in pH. By using an appropriate indicator, a chemist can tell when a solution is neutralized by monitoring its color.

Question: How is titration used to determine an unknown concentration?

Measure: Titrate the sulfuric acid analyte (H₂SO₄) with the sodium hydroxide titrant (NaOH).

How much 1.00 M NaOH is needed to neutralize the H₂SO₄ solution?

21.6 mL

Interpret: The balanced equation for the reaction of H₂SO₄ and NaOH is given at bottom right.

Based on this equation, how many moles of NaOH react with 1 mole of H₂SO₄?

2

- Manipulate: Recall that molarity is equal to the number of moles of a substance dissolved in one liter of solution: molarity = moles + volume.
 - A. Write an equation for determining the number of moles of NaOH that are added to the flask based on [NaOH] and volume of NaOH titrant (ml. NaOH):

Moles NaOH =
$$[NaOH] * mL NaOH/1,000mL$$

B. Write a similar expression for the number of moles of H₂SO₄ in the flask based on [H₂SO₄] and the volume of H₂SO₄ (mL).

Moles
$$H_2SO_4 = [H_2SO_4] * mL H_2SO_4/1,000 mL$$

C. Because there are twice as many moles of NaOH as moles of H₂SO₄ in this reaction, you can say:

Substitute your expressions from 3A and 3B into this equation and solve for [H₂SO₄]:

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Wenbin Ji

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Table of Contents Titration Virtual Activity 1 Answer Key

- 1. Understanding the eBook Titration Virtual Activity 1 Answer Key
 - The Rise of Digital Reading Titration Virtual Activity 1 Answer Key
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Titration Virtual Activity 1 Answer Key
 - Exploring Different Genres
 - o Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Titration Virtual Activity 1 Answer Key
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Titration Virtual Activity 1 Answer Key
 - Personalized Recommendations
 - Titration Virtual Activity 1 Answer Key User Reviews and Ratings
 - Titration Virtual Activity 1 Answer Key and Bestseller Lists

- 5. Accessing Titration Virtual Activity 1 Answer Key Free and Paid eBooks
 - o Titration Virtual Activity 1 Answer Key Public Domain eBooks
 - Titration Virtual Activity 1 Answer Key eBook Subscription Services
 - Titration Virtual Activity 1 Answer Key Budget-Friendly Options
- 6. Navigating Titration Virtual Activity 1 Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Titration Virtual Activity 1 Answer Key Compatibility with Devices
 - Titration Virtual Activity 1 Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Titration Virtual Activity 1 Answer Key
 - Highlighting and Note-Taking Titration Virtual Activity 1 Answer Key
 - Interactive Elements Titration Virtual Activity 1 Answer Key
- 8. Staying Engaged with Titration Virtual Activity 1 Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Titration Virtual Activity 1 Answer Key
- 9. Balancing eBooks and Physical Books Titration Virtual Activity 1 Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Titration Virtual Activity 1 Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Titration Virtual Activity 1 Answer Key
 - Setting Reading Goals Titration Virtual Activity 1 Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Titration Virtual Activity 1 Answer Key
 - Fact-Checking eBook Content of Titration Virtual Activity 1 Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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