Writing Electron Configurations

The filling order for electrons in energy sublevels is:

1s, 2s, 2p, 3s, 3p, 4s, 3d, 4p, 5s, 4d, 5p, 6s, 4f, 5d, 6p, 7s, 5f, 6d, 7p

Each s sublevel contains 1 orbital; each p contains 3; each d contains 5; and each f contains 7. Each orbital can contain at most 2 electrons.

An electron configuration can be written out by assigning electrons to the sublevels in the order listed, until the number of electrons assigned equals the atomic number (N) of the atom.

Examples

Determine the configurations of these elements.

Beryllium (N = 4) 1s²2s²

Aluminum (N = 13) $1s^22s^22p^63s^23p^1$

Bromine (N = 35) $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$

Write the electron configuration for each of the following elements.

Calcium (N = 20)

1. 15225206352306452

2. Lithium (N = 3)

152522p 353p

3. Argon (N = 18)

4 152522p6353p64523d6

Iron (N = 26)

2. 12,52,50,032,

Sodium (N = 11)

6. TS2552504

Oxygen (N = 8)

7. 15252563525645,3904 ADO 22 40, 20014574

Iodine (N = 53)

8. 15:2522p63523p64523d104p655349105p66524414594

8. Dysprosium (N = 66)

9. 125525562253 30ed2,39/10000 20020 603 04/10/29/10/00 60 4 125224,

Radium (N = 88)

Fermium (N = 100)

10. 15225206352306452391040655249105066524510591060675

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