

Thermodynamics IN-CLASS Practice Assessment:

Skill - Can define and identify direction of heat flow for exothermic and endothermic situation

- Using the image to the right, draw arrows and describe as many endothermic and exothermic processes as possible.



Skill - Use the equation $\Delta E = q + w$ correctly when given these terms in a real world situation

- What is the ΔE if a substance releases 100kJ of energy and has 52kJ of work done on it?
- How many calories of energy correspond to 28.4 J?

Skill - Use $q = sm\Delta t$ to calculate the missing quantity, given the others

- Using p329, identify the substance being described if 42.6 J is required to heat 2.0 g of this substance from 50°C to 80°C.
- Calculate the number of Calories needed to raise the temperature of 271.3 g of water from 17.1°C to 61.7°C?

Skill - Calculate the energy of a reaction using Hess's Law and identify if it is exothermic or endothermic

- Given the following data:
 $\text{CO}_2 \rightarrow \text{C} + \text{O}_2$ $\Delta H = 393.5 \text{ kJ}$
 $2 \text{ CO}_2 \rightarrow 2 \text{ CO} + \text{O}_2$ $\Delta H = 566.0 \text{ kJ}$

What is the ΔH for:



$\Delta H =$

Exothermic or

Endothermic?

- Given the following data, calculate the ΔH for the formation of ethane (C_2H_6)
 $\text{C (s)} + \text{O}_2 \text{ (g)} \rightarrow \text{CO}_2 \text{ (g)}$ $\Delta H = -393.5 \text{ kJ}$
 $\text{H}_2 \text{ (g)} + \frac{1}{2} \text{O}_2 \text{ (g)} \rightarrow \text{H}_2\text{O (l)}$ $\Delta H = -285.8 \text{ kJ}$
 $2 \text{ C}_2\text{H}_6 \text{ (g)} + 7 \text{ O}_2 \text{ (g)} \rightarrow 4 \text{ CO}_2 \text{ (g)} + 6 \text{ H}_2\text{O (l)}$ $\Delta H = -3119.6 \text{ kJ}$

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Answer Key

JR Anderson



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