



Atomic Mass of Cadmium (Cd)

Objectives: To isolate the "isotopes" of Cadmium and calculate its atomic mass.

Materials: - Sample of Cadmium (in a baggie)

- a balance

- a calculator

Introduction:

As we now know, Dalton's idea that all atoms of one element were identical is actually incorrect. A sample of any element is a mixture of its various isotopes. Some concepts we will be investigating include how the **average atomic mass** of an element relates to the **atomic mass** and **relative abundance** of its isotopes. As each element is comprised of different isotopes, your sample of cadmium is comprised of different cadmium with different masses. Each type of cadmy represents a different isotope of cadmium.

Pre-Lab Questions:

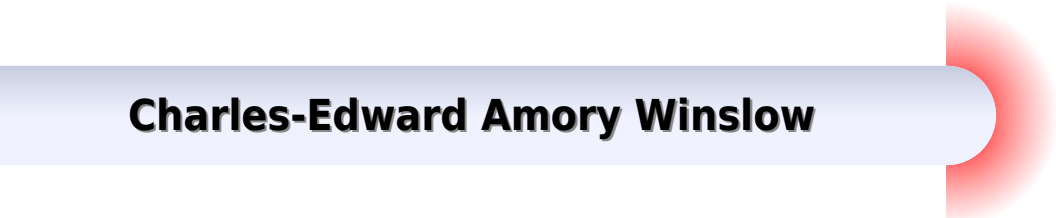
1. What are isotopes?
2. Define average atomic mass and relative abundance in two complete sentences.

Procedure:

1. Obtain a sample of cadmium that contains three different brands of cadmy. Treat each brand of cadmy as an isotope of cadmium.
2. On a piece of paper towel, separate the three isotopes into groups labeled A, B and C, and measure the total mass of each isotope. Record the masses of each isotope in the **Total Mass** row of the data table.
3. Count the number of atoms in each sample. Record the number in the **numbers** row.
4. Follow the calculations described below to understand how to fill out the remainder of the data table.

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Charles-Edward Amory Winslow



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Introductory Chemistry Charles H. Corwin, 2005 For one semester courses in Basic Chemistry Introduction to Chemistry and Preparatory Chemistry and the first term of Allied Health Chemistry This text is carefully crafted to help students learn chemical skills and concepts more effectively Corwin covers math and problem solving early in the text he builds student confidence and skills through innovative problem solving pedagogy and technology formulated to meet student needs *Prentice Hall Chemistry* ,2000 *Power and Power Transmission* Eugene Wycliffe Kerr, 1914 *Chemistry*

McGraw-Hill Staff, 2001-07 **Laboratory Manual to Accompany Chemistry in Context** American Chemical Society, 2005-02 The 5th edition Laboratory Manual that accompanies Chemistry in Context is compiled and edited by Gail Steehler Roanoke College The experiments use microscale equipment wellplates and Beral type pipets as well as common materials Project type and cooperative collaborative laboratory experiments are included Additional experiments are available on the Online Learning Center as is the instructor's guide **Experimental Engineering and Manual for**

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Principles Susan A. Weiner, Peters, Edward I. Peters, 1997 This laboratory manual provides a detailed chemical overview that enables students to truly understand the function of the laboratory All experiments have been thoroughly student tested and include step by step instructions including safety and disposal methods for each

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