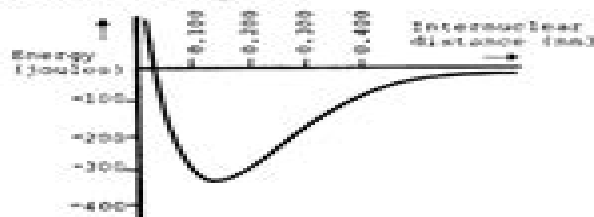


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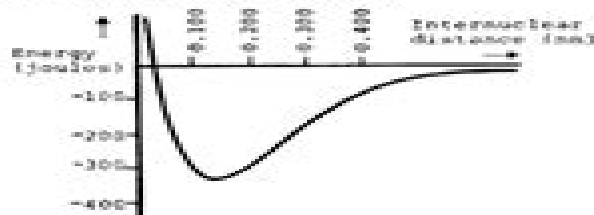
LECTURE NOTES 1

1. A bond forms when two orbitals _____ allowing electrons to be close to two nuclei at the same time. When two H atoms bond, they overlap electrons in their _____ orbitals.
2. As two H atoms approach each other, the potential energy _____ because of the electron-proton attractions. The potential energy _____ because of electron-electron and proton-proton repulsions.
3. Consider this diagram:



Two separate atoms have an energy of ____ J. The equilibrium bond distance is ____ nm. The bond energy is ____ Joules which is the amount of energy you need to add to break the bond (i.e. separate the two atoms).

4. Consider this diagram again:



If two He atoms approach, their bond distance is 0.500 nm and bond energy is 100 J. Sketch the energy curve for He₂. The two atoms do not get close enough for the electrons to be close to both nuclei at the same time... no bond forms.

We generalize: two _____ orbitals CAN make a bond.
Two _____ orbitals cannot make a bond.

5. You should be able *quickly* to state the number of valence electrons any atom has based on its position on the periodic table. Write it.

P	Ca	F	As
C	Na	K	H
Si	O	B	Mg

6. Draw the Lewis electron-dot symbol for each of the following elements.

P	Ca	F	As
C	Na	K	H
Si	O	B	Mg

7. Some elements have different "ground state" Lewis symbols. Practice drawing them:

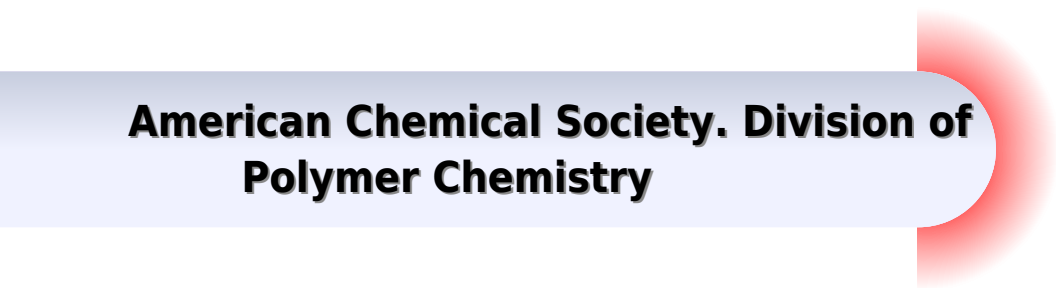
Ground State	Bonding State	Ground State	Bonding State
C	C	Ca	Ca
Al	Al	B	B

8. You should be able to recognize the **type of bond** that forms between two atoms based on the atoms involved. Classify each of the following bonds as **Ionic** or **Covalent**.

SO	_____	KI	_____
NaCl	_____	HCl	_____
CaBr	_____	SiO	_____
XeF	_____	MgCl	_____
CO	_____	NI	_____

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