

**Balancing a Redox Equation for an Acidic Solution
Using the Oxidation Number Method**
(Student textbook page 615)

35. Use the oxidation number method to balance the following equation for the combustion of carbon disulfide:



What Is Required?

You must balance an equation for the combustion of carbon disulfide, using the oxidation number method.

What Is Given?

You are given the reactants, carbon disulfide and oxygen, and the products, carbon dioxide and sulfur dioxide.

You are given rules for assigning oxidation numbers in **Table 9.3** on page 604 of the student textbook.

Plan Your Strategy	Act on Your Strategy
Step 1 Assign an oxidation number to each atom in the equation and determine whether it is a redox reaction.	$\overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g})$ <p>The oxidation numbers of sulfur and oxygen change during the reaction. It is a redox reaction.</p>
Step 2 Identify the atom or atoms that undergo an increase in oxidation number and the atom or atoms that undergo a decrease in oxidation number.	<p>Each of two sulfur atoms lost 6 electrons and each of two oxygen atoms gained 2 electrons. So the O_2 lost 12 electrons and the CS_2 gained 4 electrons.</p> $ \begin{array}{c} \text{1}(-12\text{e}^-) = -12\text{e}^- \\ \downarrow \\ \overset{+4 \ -2}{\text{CS}_2}(\text{g}) + \overset{0}{\text{O}_2}(\text{g}) \rightarrow \overset{+4 \ -2}{\text{CO}_2}(\text{g}) + \overset{+4 \ -2}{\text{SO}_2}(\text{g}) \\ \uparrow \qquad \qquad \qquad \uparrow \qquad \qquad \uparrow \\ \text{3}(+4\text{e}^-) = +12\text{e}^- \end{array} $

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Table of Contents Solutions Manual Teacher Redox

1. Understanding the eBook Solutions Manual Teacher Redox
 - The Rise of Digital Reading Solutions Manual Teacher Redox
 - Advantages of eBooks Over Traditional Books
2. Identifying Solutions Manual Teacher Redox
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Solutions Manual Teacher Redox
 - User-Friendly Interface
4. Exploring eBook Recommendations from Solutions Manual Teacher Redox
 - Personalized Recommendations
 - Solutions Manual Teacher Redox User Reviews and Ratings
 - Solutions Manual Teacher Redox and Bestseller Lists

5. Accessing Solutions Manual Teacher Redox Free and Paid eBooks
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 - Solutions Manual Teacher Redox Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Solutions Manual Teacher Redox Compatibility with Devices
 - Solutions Manual Teacher Redox Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Solutions Manual Teacher Redox
 - Highlighting and Note-Taking Solutions Manual Teacher Redox
 - Interactive Elements Solutions Manual Teacher Redox
8. Staying Engaged with Solutions Manual Teacher Redox
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Solutions Manual Teacher Redox
9. Balancing eBooks and Physical Books Solutions Manual Teacher Redox
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Solutions Manual Teacher Redox
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Solutions Manual Teacher Redox
 - Setting Reading Goals Solutions Manual Teacher Redox
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Solutions Manual Teacher Redox
 - Fact-Checking eBook Content of Solutions Manual Teacher Redox
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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