

6. Write the equilibrium expression in terms of partial pressures (K_p) for each of the following reactions.

Rate the reactions in order of their increasing tendency to proceed toward completion:

b c d a



$K_p = 1 \times 10^{28} \text{ atm}$

$$K_p = \frac{(P_{\text{N}_2})^2 (P_{\text{H}_2\text{O}})^6}{(P_{\text{NH}_3})^4 (P_{\text{O}_2})^3}$$



$K_p = 5 \times 10^{-11}$

$$K_p = \frac{(P_{\text{NO}})^2}{(P_{\text{N}_2})(P_{\text{O}_2})}$$



$K_p = 1 \times 10^{-13}$

$$K_p = \frac{(P_{\text{H}_2})(P_{\text{F}_2})}{(P_{\text{HF}})^2}$$



$K_p = 4.7 \times 10^{-4} \text{ atm}$

$$K_p = \frac{(P_{\text{NO}})^2 (P_{\text{Cl}_2})}{(P_{\text{NOCl}})^2}$$

A Question That You Should Be Able To Answer:

Why don't the K_p 's in (b) and (c) have units?

(b) $\frac{\text{atm}^2}{\text{atm}^2} = 1$ (c) $\frac{\text{atm}^2}{\text{atm}^2} = 1$

7. (a) Write the K_c expression for $2\text{SO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{SO}_3(\text{g})$

$$K_c = \frac{[\text{SO}_3]^2}{[\text{SO}_2]^2 [\text{O}_2]}$$

Calculate the value of K_c :

At equilibrium: $[\text{SO}_2] = 1.50 \text{ M}$
 $[\text{O}_2] = 1.25 \text{ M}$
 $[\text{SO}_3] = 3.50 \text{ M}$

$$= \frac{(3.50)^2}{(1.50)^2 (1.25)} = 4.355$$

$$= 4.36 \text{ M}^{-1}$$

- (b) If we reverse the equation, it is:



Write the K_c expression for this equation and calculate the new value of K_c :

$$K_c' = \frac{[\text{SO}_2]^2 [\text{O}_2]}{[\text{SO}_3]^2} = \frac{(1.5)^2 (1.25)}{(3.50)^2} = .22959$$

How does the expression and the value of K_c in 7(b) compare with those in 7(a)?

$$K_c' = \frac{1}{K_c}$$

- (c) If we now multiply all of the coefficients by $\frac{1}{2}$:



Write the K_c expression for this equation and calculate the new value of K_c :

$$K_c'' = \frac{[\text{SO}_2] [\text{O}_2]^{1/2}}{[\text{SO}_3]} = \frac{(1.5) (1.25)^{1/2}}{(3.50)} = .479$$

How do they compare with 7(b)?

$$K_c'' = \sqrt{K_c'}$$

- (d) What would happen to the K_c expression and its value if we doubled the coefficients?

$$K_c''' = (K_c')^2$$

Summarize:

Equation	K_c expression & Value
doubled	SQUARE
reversed	inverse
halved	SQ. ROOT

8. Consider an equilibrium that occurs in two steps:



- (a) Write the overall reaction.

(b) How do the K_c 's for the two steps (K_{c1} & K_{c2}) relate to the K_c of the overall reaction (K_c)?

$$K_{c1} \cdot K_{c2} = K_c$$

South Pasadena Chemistry Answers

Camilla Rothe



South Pasadena Chemistry Answers:

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web the upper secondary science normal technical syllabus is based on the science curriculum framework the syllabus is for a two year course that is designed to provide students with an in this respect the development of scientific literacy supports moe s efforts on the development of students 21cc as discussed in section 1 1

subjects and syllabuses moe - Oct 05 2022

web feb 4 2021 primary school subjects and syllabuses primary school students will be introduced to subject based learning where they will learn subjects such as languages mathematics science art music and social studies learn more about the syllabus taught in the classroom