

# Solution

## Section A

- (d) Molality (m)

**Explanation:** Molality is defined as no. of moles of solute present per kg of solvent. It is denoted as " m ". Mathematically, Molality (m) = [number of moles of the solute] / kg of solvent  
Since this mode of expressing the strength of a solution involves (weight/weight) relationship of solute and solvent, the molality of the solution is not affected by variation in temperatures of the solution.
- (a) behaves like particles

**Explanation:** The emission of free electrons from a metal surface when light is shone on it is called the photoemission or the photoelectric effect. This effect led to the conclusion that light is made up of packets or quantum of energy. Einstein already associated the light quantum with momentum. This strongly supported the particle nature of light and these particles were named photons. Thus, the wave-particle duality of light came into the picture. Einstein won the Nobel Prize for Physics not for his work on relativity, but for explaining the photoelectric effect.
- (b)  $\Delta H = -7.151 \text{ kJ mol}^{-1}$

**Explanation:** Water (10.0°C)  $\rightarrow$  ice (-10.0°C)  $\Delta H = ?$   
The enthalpy change for the conversion of 1 mole liquid water at 10.0°C into 1 mole liquid water 0°C, water (10.0°C)  $\rightleftharpoons$  water(0°C).  
 $\Delta H_1 = C_p \times H_2O(l) \times \Delta T = -75.3 \text{ Jmol}^{-1} \text{ K}^{-1} = -753 \text{ J mol}^{-1}$   
Enthalpy of fusion, water (0°C)  $\rightleftharpoons$  ice(0°C),  $\Delta H_2 = \Delta_{fus}H$   
 $\Delta H_2 = \Delta H_{fusion} = -\Delta H_{freezing} = -6.06 \text{ kJ mol}^{-1}$   
Enthalpy change for conversion of 1 mole of ice at 0°C to 1 mole of ice at 10°C, Ice (0°C)  $\rightleftharpoons$  ice(10.0°C);  
 $\Delta H_3 = C_p[H_2O(s)]\Delta T = 36.8 \text{ Jmol}^{-1}\text{K}^{-1} \times 10\text{K} = 368 \text{ J mol}^{-1}$   
 $\Delta H = \Delta H_1 + \Delta H_2 + \Delta H_3 = -(0.753 + 6.03 + 0.368)\text{kJ mol}^{-1} = -7.151 \text{ kJmol}^{-1}$
- (d)  $1.988 \times 10^{-18} \text{ J}$

**Explanation:** We know Planck's equation is  
 $E = h\nu$   
where E is energy, h is Planck's constant and  $\nu$  is frequency.  
 $E = 6.626 \times 10^{-34} \times 3 \times 10^{15} = 1.988 \times 10^{-18} \text{ J}$
- (c) 20.15

**Explanation:**  $\Delta G^\circ = -115 \times 10^3 \text{ J}$   
 $T = 298 \text{ K}$ ,  $R = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$   
 $-\Delta G^\circ = 2.303 RT \log_{10} K_p$   
 $-(-115 \times 10^3) = 2.303 \times 8.314 \times 298 \log_{10} K_p$   
 $\log_{10} K_p = \frac{115000}{2.303 \times 8.314 \times 298} = 20.15$
- (d) Paschen series

**Explanation:** The Lyman series lies in the ultraviolet, whereas the Paschen, Brackett, and Pfund series lies in the infrared.
- (b) reduction

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