

SS 2 CHEMISTRY PRACTICAL 3RD TERM EXMINATION

1. B is a solution containing 5.30g of anhydrous sodium trioxocarbonate (IV) per dm³. A is a solution of hydrochloric acid of unknown concentration.
Put A into the burette and titrate 25cm³ (or 20cm³) portions of B with A. Using methyl orange as the indicator. Record the volume of your pipette.
Tabulate your burette readings, and calculate the average volume of acid used.
From your results, calculate:
 - a. The concentration of solution B in moles per dm³
 - b. The concentration of solution A in moles per dm³
 - c. The mass concentration in g per dm³ of the acid in solution A

The equation of reaction is :



(H = 1; C = 12; O = 16; Na = 23; Cl = 35.5)

2. (a). If you were provided with anhydrous CaCO₃, spatula and stirrer (i) list three other materials you would require to prepare a standard solution of CaCO₃ (ii) State what you would observe on adding dilute HCl to a portion of the CaCO₃
(b). State what would be observed when aqueous ammonia is added to: (i) Litmus paper; (ii) Pb(NO₃)₂ solution in drops until in excess.

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