



LAGOS STATE GOVERNMENT
MINISTRY OF BASIC AND SECONDARY EDUCATION
3RD TERM UNIFIED EXAMINATION 2024

SS 1

CHEMISTRY

Time: 2 hrs.

SECTION A - OBJECTIVES

INSTRUCTION: Answer all questions in this section.

- The branch of science that deals with the nature and properties of substances and how one substance can be converted to another is known as
(a) Biology (b) Geography
(c) Chemistry (d) Physics
- The electron configuration $1s^2 2s^2 2p_x^2 2p_y^2$ depicts the
(a) Pauli's exclusion principle
(b) Aufbau's principle
(c) Octet rule
(d) Hund's rule
- The oxidation number of chromium in $Na_2Cr_2O_7$ is
(a) +2 (b) +6
(c) +7 (d) +12
- Bases normally
(a) are corrosive
(b) turn litmus paper from blue to red
(c) turn litmus paper from red to blue
(d) are non-metal oxides
- Naturally occurring Boron is made up of 19.9% ^{10}B and 80.1% ^{11}B . The relative atomic mass of Boron is
(a) 21.0 (b) 10.8
(c) 10.5 (d) 10.0
$$\frac{19.9 \times 10 + 80.1 \times 11}{100} = 10.81$$
- Isotopes of the same element have similar chemical properties because they have the same number of
(a) Nuclides (b) Neutrons
(c) Atoms (d) protons
- Which of the following household liquids would be effective in treating someone with too much acid in the stomach?
(a) Strong salt solution with pH 7.0
(b) Tomato juice with pH 4.1
(c) Milk of Magnesia with pH 10.5
(d) Black coffee with pH 5.0
- The percentage by mass of oxygen in $MgSO_4 \cdot 7H_2O$ is
(a) 26.0% (b) 45.5%
(c) 71.5% (d) 84.0%
- Ethane of volume $400cm^3$ was completely burnt in excess oxygen according to the following equation:
 $2C_2H_6(g) + 7O_2(g) \longrightarrow 6H_2O_{(g)} + 4CO_{2(g)}$
Calculate the volume of steam that would be produced.
(a) $200cm^3$ (b) $400cm^3$
(c) $600cm^3$ (d) $1200cm^3$
$$\frac{2 \times 400}{2} = 400 \quad \frac{6 \times 1200}{4} = 1800$$

$$400 = x \quad 1800 = 1200$$
- A weak acid is one which
(a) is not corrosive
(b) completely ionizes in water
(c) does not produce salt with alkali
(d) slightly ionizes in water

Sss1 Chemistry Outline

**Achim A. Beylich, John C.
Dixon, Zbigniew Zwoliński**



Sss1 Chemistry Outline:

The book delves into Sss1 Chemistry Outline. Sss1 Chemistry Outline is a crucial topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Sss1 Chemistry Outline, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:

- Chapter 1: Introduction to Sss1 Chemistry Outline
- Chapter 2: Essential Elements of Sss1 Chemistry Outline
- Chapter 3: Sss1 Chemistry Outline in Everyday Life
- Chapter 4: Sss1 Chemistry Outline in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, the author will provide an overview of Sss1 Chemistry Outline. The first chapter will explore what Sss1 Chemistry Outline is, why Sss1 Chemistry Outline is vital, and how to effectively learn about Sss1 Chemistry Outline.
3. In chapter 2, the author will delve into the foundational concepts of Sss1 Chemistry Outline. The second chapter will elucidate the essential principles that must be understood to grasp Sss1 Chemistry Outline in its entirety.
4. In chapter 3, the author will examine the practical applications of Sss1 Chemistry Outline in daily life. The third chapter will showcase real-world examples of how Sss1 Chemistry Outline can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Sss1 Chemistry Outline in specific contexts. This chapter will explore how Sss1 Chemistry Outline is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Sss1 Chemistry Outline. The final chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of Sss1 Chemistry Outline.

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