

Reporting Category 1: Matter and the Periodic Table (12 questions)

C.4A	RS	differentiate between physical and chemical changes and properties
C.4B	SS	identify extensive and intensive properties
C.4C	SS	compare solids, liquids, and gases in terms of compressibility, structure, shape, and volume
C.4D	RS	classify matter as pure substances or mixtures through investigation of their properties
C.5A	SS	explain the use of chemical and physical properties in the historical development of the Periodic Table
C.5B	RS	use the Periodic Table to identify and explain the properties of chemical families, including alkali metals, alkaline earth metals, halogens, noble gases, and transition metals
C.5C	RS	use the Periodic Table to identify and explain periodic trends, including atomic and ionic radii, electronegativity, and ionization energy

Reporting Category 2: Atomic Structure and Nuclear Chemistry (9 questions)

C.6A	SS	understand the experimental design and conclusions used in the development of modern atomic theory, including Dalton's Postulates, Thomson's discovery of electron properties, Rutherford's nuclear atom, and Bohr's nuclear atom
C.6B	SS	understand the electromagnetic spectrum and the mathematical relationships between energy, frequency, and wavelength of light
C.6C	SS	calculate the wavelength, frequency, and energy of light using Planck's constant and the speed of light
C.6D	SS	use isotopic composition to calculate average atomic mass of an element
C.6E	RS	express the arrangement of electrons in atoms through electron configurations and Lewis valence electron dot structures
C.12A	SS	describe the characteristics of alpha, beta, and gamma radiation
C.12B	RS	describe radioactive decay process in terms of balanced nuclear equations
C.12C	SS	compare fission and fusion reactions

Reporting Category 3: Bonding and Chemical Reactions (14 questions)

C.7A	RS	name ionic compounds containing main group or transition metals, covalent compounds, acids, and bases, using International Union of Pure and Applied Chemistry (IUPAC) nomenclature rules
C.7B	RS	write the chemical formulas of common polyatomic ions, ionic compounds containing main group or transition metals, covalent compounds, acids, and bases
C.7C	RS	construct electron dot formulas to illustrate ionic and covalent bonds
C.7D	SS	describe the nature of metallic bonding and apply the theory to explain metallic properties such as thermal and electrical conductivity, malleability, and ductility
C.7E	SS	predict molecular structure for molecules with linear, trigonal planar, or tetrahedral electron pair geometries using Valence Shell Electron Pair Repulsion (VSEPR) theory
C.8A	SS	define and use the concept of a mole
C.8B	RS	use the mole concept to calculate the number of atoms, ions, or molecules in a sample of material
C.8C	SS	calculate percent composition and empirical and molecular formulas
C.8D	RS	use the law of conservation of mass to write and balance chemical equations
C.8E	SS	perform stoichiometric calculations, including determination of mass relationships between reactants and products, calculation of limiting reagents, and percent yield

Reporting Category 4: Gases and Thermochemistry (8 questions)

C.9A	RS	describe and calculate the relations between volume, pressure, number of moles, and temperature for an ideal gas as described by Boyle's law, Charles' law, Avogadro's law, Dalton's law of partial pressure, and the ideal gas law
C.9B	SS	perform stoichiometric calculations, including determination of mass and volume relationships between reactants and products for reactions involving gases
C.9C	SS	describe the postulates of kinetic molecular theory
C.11A	SS	understand energy and its forms, including kinetic, potential, chemical, and thermal energies
C.11B	SS	understand the law of conservation of energy and the processes of heat transfer
C.11C	RS	use thermochemical equations to calculate energy changes that occur in chemical reactions and classify reactions as exothermic or endothermic
C.11D	SS	perform calculations involving heat, mass, temperature changes, and specific heat
C.11E	SS	use calorimetry to calculate the heat of a chemical process

Staar Chemistry Quick Reference Guide

R Bogdan



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Reference Sources Linda Mark,1979 *Chemistry Quick Reference Guide* Innovative Academic Solutions,2007

The New York Times Book Review ,1964 *ACRL News* ,1969 **Books in Print** ,1991 **College & Research**

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