

STAAR Science Tutorial 06

TEK 8.5B: Atomic Structure & Chemical Properties

TEK 8.5B: *Identify that protons determine an element's identity and valence electrons determine its chemical properties, including reactivity.*

Valence Electrons & Chemical Properties

- Each element has a unique number of protons in all of its atoms. The **atomic number** of an element is the number of protons in each of its atoms. To identify which element a model of an atom is depicting, count the number of protons in the nucleus and find the element on the periodic table with that atomic number.
- The total number of electrons in an (electrically neutral) atom equals the number of protons. Thus, an atom can also be identified by the total number of electrons shown in the Electron Cloud Model or Bohr Model.
- In the Bohr Model of the atom, electrons fill different electron orbitals (also known as electron shells or energy levels) from the inside-out. The first orbital can only hold 2 electrons. The second orbital can hold 8 electrons, the third can hold up to 18, and the fourth up to 32, but the outermost orbital can never have more than 8 electrons. Only these outermost electrons, known as **valence electrons**, can interact with other atoms in chemical reactions.
- The following chart summarizes the maximum number of electrons that can be in each orbital (if it is not the outermost orbital):

Orbital Number	1	2	3	4	5	6	7
No. of Electrons	2	8	18	32	32	18	8

- In general, elements with the same number of valence electrons, usually found in the same group on the periodic table, have similar chemical properties.
- For the Representative Elements in Groups 1 - 2 and 13 -18, the number of valence electrons is equal to the ones digit of the group number. For example, there is 1 valence electron in Group 1 elements, and 3 valence electrons in Group 13 elements.
- The following chart summarizes the number of valence electrons for the representative elements:

Group/Family Number	1	2	13	14	15	16	17	18
Valence Electrons	1	2	3	4	5	6	7	8

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Karin Nielsen-Saines



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Table of Contents Staar Science Tutorial 06

1. Understanding the eBook Staar Science Tutorial 06
 - The Rise of Digital Reading Staar Science Tutorial 06
 - Advantages of eBooks Over Traditional Books
2. Identifying Staar Science Tutorial 06
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Staar Science Tutorial 06
 - User-Friendly Interface
4. Exploring eBook Recommendations from Staar Science Tutorial 06
 - Personalized Recommendations
 - Staar Science Tutorial 06 User Reviews and Ratings
 - Staar Science Tutorial 06 and Bestseller Lists
5. Accessing Staar Science Tutorial 06 Free and Paid eBooks
 - Staar Science Tutorial 06 Public Domain eBooks
 - Staar Science Tutorial 06 eBook Subscription Services
 - Staar Science Tutorial 06 Budget-Friendly Options
6. Navigating Staar Science Tutorial 06 eBook Formats

- ePub, PDF, MOBI, and More
- Staar Science Tutorial 06 Compatibility with Devices
- Staar Science Tutorial 06 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Staar Science Tutorial 06
 - Highlighting and Note-Taking Staar Science Tutorial 06
 - Interactive Elements Staar Science Tutorial 06
- 8. Staying Engaged with Staar Science Tutorial 06
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Staar Science Tutorial 06
- 9. Balancing eBooks and Physical Books Staar Science Tutorial 06
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Staar Science Tutorial 06
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Staar Science Tutorial 06
 - Setting Reading Goals Staar Science Tutorial 06
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Staar Science Tutorial 06
 - Fact-Checking eBook Content of Staar Science Tutorial 06
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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